

College Prep Chemistry of the Earth System

Assignment 2D

30 Points

Specific Heat Calculations

Solve the following specific heat problems

1. $c_{\text{water}} = 4.186 \text{ J/g}^\circ\text{C}$, $m = 3.62 \text{ g}$,
 $\Delta T = 15.2^\circ\text{C}$, $q = \underline{\hspace{2cm}} \text{ J}$
2. $c_{\text{ice}} = 2.090 \text{ J/g}^\circ\text{C}$, $m = 84.6 \text{ g}$,
 $\Delta T = 95.1^\circ\text{C}$, $q = \underline{\hspace{2cm}} \text{ J}$
3. $q = 934.5 \text{ J}$, $m = 42.5 \text{ g}$
 $\Delta T = 29.2^\circ\text{C}$, $c = \underline{\hspace{2cm}} \text{ J/g}^\circ\text{C}$
4. $q = 254.9 \text{ J}$, $m = 28.8 \text{ g}$
 $\Delta T = 15.0^\circ\text{C}$, $c = \underline{\hspace{2cm}} \text{ J/g}^\circ\text{C}$
5. $q = 1943.5 \text{ J}$, $c = 0.482 \text{ J/g}^\circ\text{C}$
 $\Delta T = 35.2^\circ\text{C}$, $m = \underline{\hspace{2cm}} \text{ g}$
6. $q = 835.8 \text{ J}$, $c = 0.284 \text{ J/g}^\circ\text{C}$
 $\Delta T = 21.6^\circ\text{C}$, $m = \underline{\hspace{2cm}} \text{ g}$
7. $q = 913.5 \text{ J}$, $c = 0.834 \text{ J/g}^\circ\text{C}$
 $m = 29.4 \text{ g}$, $\Delta T = \underline{\hspace{2cm}}^\circ\text{C}$
8. $q = 2184.2 \text{ J}$, $c = 0.184 \text{ J/g}^\circ\text{C}$
 $m = 19.3 \text{ g}$, $\Delta T = \underline{\hspace{2cm}}^\circ\text{C}$