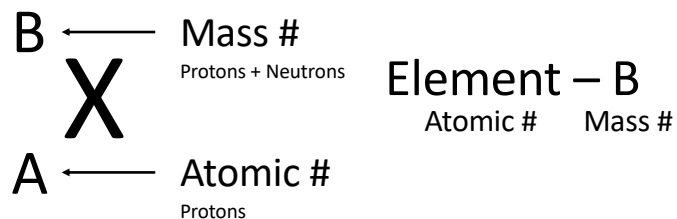


Atomic Isotope Notation

Writing atomic isotopes



2

Compare Mass Number and Atomic Mass

Mass Number

The number of subatomic particles that make up the mass of an atom ($p^+ + n^0$, count)

Example

Sodium – 23 ($11p^+ + 12n^0$)
 Mass Number = $11 + 12 = 23$

Atomic Mass

The mass of the subatomic particles that contribute mass to an atom ($p^+ + n^0$, amu)

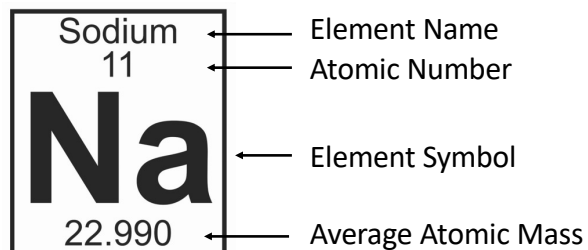
Example

Sodium – 23 ($11p^+ + 12n^0$)
 Atomic Mass = $11\text{amu} + 12\text{amu}$
 = 23amu

3

Average Atomic Mass

The average of the atomic masses of all isotopes that make up a particular form of matter.



4

Calculating Average Atomic Mass

Average Atomic Mass is calculated based on the equation below

$$\sum(\text{Atomic Mass} \times \text{Frac}_{\text{Abundance}})$$

$$\text{Frac}_{\text{Abundance}} = \frac{\% \text{Abundance}}{100\%} \quad \Sigma = \text{Sum (add)}$$

5