

College Prep Chemistry of the Earth System

Assignment 5E 30 Pts Total

Single Replacement w/Polyatomic Balancing

For the following chemical equations balance the equation. Show balancing chart.

- $1. \quad \underline{\quad} \text{Na} + \underline{\quad} \text{Al}(\text{NO}_3)_3 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Al} + \underline{\quad} \text{NaNO}_3$
- $2. \quad \underline{\quad} \text{Cu} + \underline{\quad} \text{AgClO}_3 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Ag} + \underline{\quad} \text{Cu}(\text{ClO}_3)_2$
- $3. \quad \underline{\quad} \text{Zn} + \underline{\quad} \text{NiSO}_4 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Ni} + \underline{\quad} \text{Zn}_2\text{SO}_4$
- $4. \quad \underline{\quad} \text{Pb} + \underline{\quad} \text{Hg}_3\text{PO}_4 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Hg} + \underline{\quad} \text{Pb}_3(\text{PO}_4)_4$
- $5. \quad \underline{\quad} \text{Mg} + \underline{\quad} \text{Sb}(\text{CrO}_4)_2 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Sb} + \underline{\quad} \text{MgCrO}_4$
- $6. \quad \underline{\quad} \text{Li} + \underline{\quad} \text{Na}_2\text{CO}_3 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Na} + \underline{\quad} \text{Li}_2\text{CO}_3$
- $7. \quad \underline{\quad} \text{Al} + \underline{\quad} \text{Sn}(\text{OH})_4 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Sn} + \underline{\quad} \text{Al}(\text{OH})_3$
- $8. \quad \underline{\quad} \text{Cr} + \underline{\quad} \text{Cu}(\text{BrO}_3)_3 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Cu} + \underline{\quad} \text{Cr}(\text{BrO}_3)_2$
- $9. \quad \underline{\quad} \text{Na} + \underline{\quad} \text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Ca} + \underline{\quad} \text{NaC}_2\text{H}_3\text{O}_2$
- $10. \quad \underline{\quad} \text{Ni} + \underline{\quad} \text{Pd}_2(\text{C}_2\text{O}_3)_3 \rightarrow$
 $\quad \quad \quad \underline{\quad} \text{Pd} + \underline{\quad} \text{NiC}_2\text{O}_3$