

College Prep Chemistry of the Earth System

Assignment 5G 50 Pts Total

Chemical Equation Balancing Review

For the following chemical equations balance the equation. Show balancing chart.

1. $\underline{\quad} \text{Ni} + \underline{\quad} \text{Br}_2 \rightarrow \underline{\quad} \text{NiBr}_3$
2. $\underline{\quad} \text{Pb} + \underline{\quad} \text{N}_2 \rightarrow \underline{\quad} \text{Pb}_3\text{N}_4$
3. $\underline{\quad} \text{H}_3\text{P} \rightarrow \underline{\quad} \text{H}_2 + \underline{\quad} \text{P}$
4. $\underline{\quad} \text{SO}_3 \rightarrow \underline{\quad} \text{S} + \underline{\quad} \text{O}_2$
5. $\underline{\quad} \text{Ca} + \underline{\quad} \text{AlCl}_3 \rightarrow$
 $\underline{\quad} \text{Al} + \underline{\quad} \text{CaCl}_2$
6. $\underline{\quad} \text{Pb} + \underline{\quad} \text{Ag}_2\text{SO}_4 \rightarrow$
 $\underline{\quad} \text{Ag} + \underline{\quad} \text{Pb}(\text{SO}_4)_2$
7. $\underline{\quad} \text{Mg} + \underline{\quad} \text{Co}(\text{OH})_3 \rightarrow$
 $\underline{\quad} \text{Co} + \underline{\quad} \text{Mg}(\text{OH})_2$
8. $\underline{\quad} \text{VBr}_3 + \underline{\quad} \text{TiCl}_4 \rightarrow$
 $\underline{\quad} \text{VCl}_3 + \underline{\quad} \text{TiBr}_4$
9. $\underline{\quad} \text{PdN} + \underline{\quad} \text{NiF}_2 \rightarrow$
 $\underline{\quad} \text{PdF}_3 + \underline{\quad} \text{Ni}_3\text{N}_2$
10. $\underline{\quad} \text{Sr}(\text{NO}_3)_2 + \underline{\quad} \text{Zr}_2\text{O} \rightarrow$
 $\underline{\quad} \text{SrO} + \underline{\quad} \text{ZrNO}_3$
11. $\underline{\quad} \text{Y}_2(\text{CO}_3)_3 + \underline{\quad} \text{CaS} \rightarrow$
 $\underline{\quad} \text{Y}_2\text{S}_3 + \underline{\quad} \text{CaCO}_3$
12. $\underline{\quad} \text{Ti}(\text{CrO}_4)_2 + \underline{\quad} \text{Cd}(\text{NO}_3)_2 \rightarrow$
 $\underline{\quad} \text{Ti}(\text{NO}_3)_4 + \underline{\quad} \text{CdCrO}_4$
13. $\underline{\quad} \text{MnPO}_4 + \underline{\quad} \text{CuSO}_4 \rightarrow$
 $\underline{\quad} \text{Mn}_2(\text{SO}_4)_3 + \underline{\quad} \text{Cu}_3(\text{PO}_4)_2$
14. $\underline{\quad} \text{HClO}_3 + \underline{\quad} \text{Cu}(\text{OH})_3 \rightarrow$
 $\underline{\quad} \text{Cu}(\text{ClO}_3)_3 + \underline{\quad} \text{HOH}$
15. $\underline{\quad} \text{H}_2\text{SO}_4 + \underline{\quad} \text{Sn}(\text{OH})_4 \rightarrow$
 $\underline{\quad} \text{Sn}(\text{SO}_4)_2 + \underline{\quad} \text{HOH}$