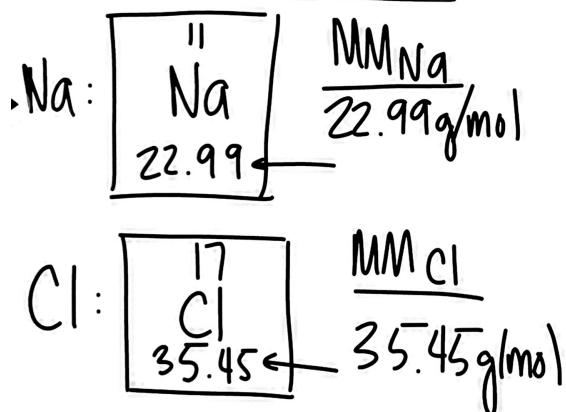


Molar Mass : The mass of 1 mol of any substance (atom, ion, comp., molecule)

Avg. Molar Mass : MM of all isotopes of element.

<u>Molar Mass</u>	<u>Avg. Molar Mass</u>						
$^{12}_6\text{C}, 12\text{g}$ $^{13}_6\text{C}, 13\text{g}$ $(\text{MM} = \text{g/mol})$	<table border="1"> <tr> <td>6</td><td>← Atomic #</td></tr> <tr> <td>C</td><td>← Element</td></tr> <tr> <td>Carbon 12.01</td><td>← Avg. MM</td></tr> </table>	6	← Atomic #	C	← Element	Carbon 12.01	← Avg. MM
6	← Atomic #						
C	← Element						
Carbon 12.01	← Avg. MM						

Combined Molar Mass



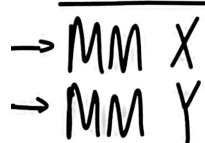
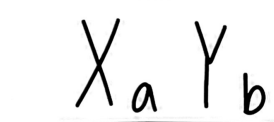
MM NaCl

MM Na + MM Cl

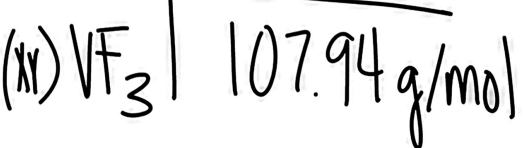
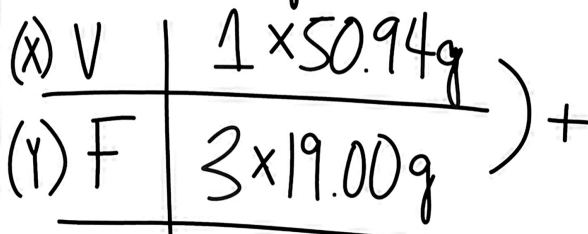
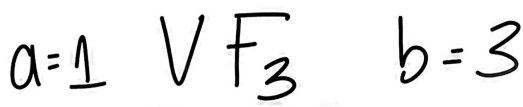
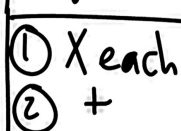
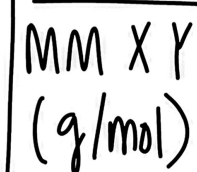
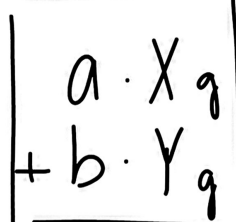
<u>MM NaCl</u>		① X each
		② +
Na	1 X 22.99g	)
Cl	1 X 35.45g	
NaCl	<u>58.44 g/mol</u>	

Molar Mass Format

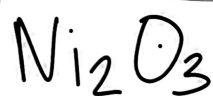
Mar-15-2023 9:40 PM 3/5



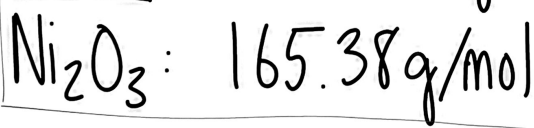
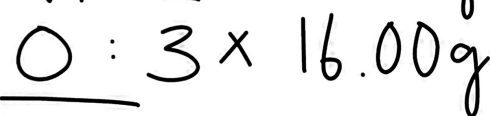
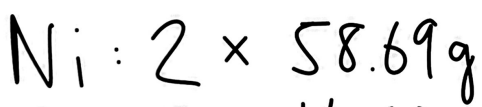
a and b are
subscripts



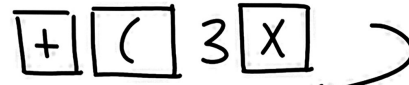
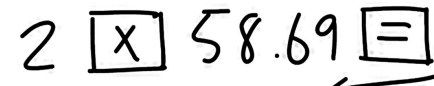
← → ↺ ↻



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Calculator



← → ↺ ↻



$$[MM_{Ti} = 47.87g][MM_S = 32.07g]$$

$$Ti: 1 \times 47.87g \quad S: 2 \times 32.07g$$

$$TiS_2: \underline{112.01g/mol}$$