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1a: Na_3N (a = 3, b = 1)

Na: $3 \times 22.99\,g$

N: $1 \times 14.01\,g$

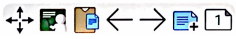
Na_3N: $82.98\,g/mol$ (Na_3N)

2a: $452.3\,g\ Na_3N \rightarrow mol$

$$452.3\,g\,Na_3N \left| \frac{1\ mol\ Na_3N}{82.98\,g\,Na_3N} \right. = 5.45\,mol\ Na_3N$$

2e: $4.29\,mol\ Na_3N \rightarrow mass$

$$4.29\,mol\,Na_3N \left| \frac{82.98\,g\ Na_3N}{1\ mol\,Na_3N} \right. = 355.98\,g\ Na_3N$$



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1b: Na_3PO_4

Na: $3 \times 22.99\,g$

P: $1 \times 30.97\,g$

O: $4 \times 16.00\,g$

Na_3PO_4: $163.94\,g/mol$ Na_3PO_4

2b: $592.5\,g\ Na_3PO_4 \rightarrow mol$

$$592.5\,g\,Na_3PO_4 \left| \frac{1\ mol\ Na_3PO_4}{163.94\,g\,Na_3PO_4} \right. = 3.61\,mol\ Na_3PO_4$$

2e: $1.49\,mol\ Na_3PO_4 \rightarrow mass$

$$1.49\,mol\,Na_3PO_4 \left| \frac{163.94\,g\ Na_3PO_4}{1\ mol\,Na_3PO_4} \right. = 244.27\,g\ Na_3PO_4$$