

Name KeyPeriod All**Quiz 19** – Combination and Single Replacement Reaction Balancing

20 Points

Balance the following chemical reactions, showing balancing chart. No Work = No Credit!

Combination Reaction	Single Replacement Reaction
2 each (6) $\underline{3} \text{ Fe} + \underline{\quad} \text{ N}_2 \rightarrow \underline{\quad} \text{ Fe}_3\text{N}_2$	$\underline{6} \text{ K} + \underline{\quad} \text{ Cu}_2\text{O}_3 \rightarrow$ 2 each (8) $\underline{2} \text{ Cu} + \underline{3} \text{ K}_2\text{O}$
Balancing Chart $\text{Fe} - \cancel{3} \quad \text{Fe} - 3$ $\text{N} - 2 \quad \text{N} - 2$ <u>work 4</u>	Balancing Chart $\text{K} - \cancel{6} \quad \text{K} - 2 \cdot 6$ $\text{Cu} - 2 \quad \text{Cu} - \cancel{2}$ $\text{O} - 3 \quad \text{O} - \cancel{3}$ <u>work 2</u>

Name KeyPeriod All**Quiz 19** – Combination and Single Replacement Reaction Balancing

20 Points

Balance the following chemical reactions, showing balancing chart. No Work = No Credit!

Combination Reaction	Single Replacement Reaction
2 each (6) $\underline{2} \text{ Al} + \underline{3} \text{ S} \rightarrow \underline{\quad} \text{ Al}_2\text{S}_3$	$\underline{3} \text{ Ca} + \underline{2} \text{ Zn}_3\text{P} \rightarrow$ 2 each (8) $\underline{6} \text{ Zn} + \underline{\quad} \text{ Ca}_3\text{P}_2$
Balancing Chart $\text{Al} - \cancel{2} \quad \text{Al} - 2$ $\text{S} - \cancel{3} \quad \text{S} - 3$ <u>work 4</u>	Balancing Chart $\text{Ca} - \cancel{3} \quad \text{Ca} - 3$ $\text{Zn} - \cancel{6} \quad \text{Zn} - \cancel{6}$ $\text{P} - \cancel{2} \quad \text{P} - 2$ <u>work 2</u>

Name key Period All

Quiz 19 - Combination and Single Replacement Reaction Balancing 20 Points
 Balance the following chemical reactions, showing balancing chart. No Work = No Credit!

Combination Reaction	Single Replacement Reaction
$\overset{2 \text{ each } (6)}{4} \text{V} + \overset{3}{3} \text{O}_2 \rightarrow \overset{2}{2} \text{V}_2\text{O}_3$	$\overset{2 \text{ each } (8)}{4} \text{Ag} + \text{PbS}_2 \rightarrow \text{Pb} + \overset{2}{2} \text{Ag}_2\text{S}$
<p style="text-align: center;"><i>Balancing Chart</i></p> $\text{V} - \cancel{4} \quad \text{V} - 2 \cdot 4$ $\text{O} - \cancel{6} \quad \text{O} - \cancel{6}$ <p style="text-align: center;"><u>work 4</u></p>	<p style="text-align: center;"><i>Balancing Chart</i></p> $\text{Ag} - \cancel{4} \quad \text{Ag} - 2 \cdot 4$ $\text{Pb} - 1 \quad \text{Pb} - 1$ $\text{S} - 2 \quad \text{S} - \cancel{2}$ <p style="text-align: center;"><u>work 2</u></p>

Name key Period All

Quiz 19 - Combination and Single Replacement Reaction Balancing 20 Points
 Balance the following chemical reactions, showing balancing chart. No Work = No Credit!

Combination Reaction	Single Replacement Reaction
$\overset{2 \text{ each } (6)}{3} \text{Li} + \text{P} \rightarrow \text{Li}_3\text{P}$	$\overset{2 \text{ each } (8)}{3} \text{Mg} + \overset{2}{2} \text{VN} \rightarrow \text{V} + \text{Mg}_3\text{N}_2$
<p style="text-align: center;"><i>Balancing Chart</i></p> $\text{Li} - \cancel{3} \quad \text{Li} - 3$ $\text{P} - 1 \quad \text{P} - 1$ <p style="text-align: center;"><u>work 4</u></p>	<p style="text-align: center;"><i>Balancing Chart</i></p> $\text{Mg} - \cancel{3} \quad \text{Mg} - 3$ $\text{V} - \cancel{2} \quad \text{V} - \cancel{2}$ $\text{N} - \cancel{2} \quad \text{N} - 2$ <p style="text-align: center;"><u>work 2</u></p>

Name keyPeriod All

Quiz 20 - Polyatomic Ion Reaction Balancing

20 Points

Balance the following chemical reactions, showing balancing chart. No Work = No Credit!

Single Replacement Reaction	Double Replacement Reaction														
$\underline{3} \text{Mn} + \underline{2} \text{Fe(OH)}_3 \rightarrow$ $\underline{2} \text{Fe} + \underline{3} \text{Mn(OH)}_2$	$\text{PbCl}_4 + \underline{2} \text{CdC}_2\text{O}_4 \rightarrow$ $\text{Pb(C}_2\text{O}_4)_2 + \underline{2} \text{CdCl}_2$														
<p>Balancing Chart</p> <table style="border: none;"> <tr> <td>Mn: 13</td> <td>Mn: 13</td> </tr> <tr> <td>Fe: 12</td> <td>Fe: 12</td> </tr> <tr> <td>OH: 36</td> <td>OH: 36</td> </tr> </table> <p style="text-align: center;">work = 2</p>	Mn: 1 3	Mn: 1 3	Fe: 1 2	Fe: 1 2	OH: 3 6	OH: 3 6	<p>Balancing Chart</p> <table style="border: none;"> <tr> <td>Pb: 1</td> <td>Pb: 1</td> </tr> <tr> <td>Cl: 4</td> <td>Cl: 24</td> </tr> <tr> <td>Cd: 12</td> <td>Cd: 12</td> </tr> <tr> <td>C₂O₄: 12</td> <td>C₂O₄: 2</td> </tr> </table> <p style="text-align: center;">work = 2</p>	Pb: 1	Pb: 1	Cl: 4	Cl: 2 4	Cd: 1 2	Cd: 1 2	C ₂ O ₄ : 1 2	C ₂ O ₄ : 2
Mn: 1 3	Mn: 1 3														
Fe: 1 2	Fe: 1 2														
OH: 3 6	OH: 3 6														
Pb: 1	Pb: 1														
Cl: 4	Cl: 2 4														
Cd: 1 2	Cd: 1 2														
C ₂ O ₄ : 1 2	C ₂ O ₄ : 2														

Name keyPeriod All

Quiz 20 - Polyatomic Ion Reaction Balancing

20 Points

Balance the following chemical reactions, showing balancing chart. No Work = No Credit!

Single Replacement Reaction	Double Replacement Reaction														
$\underline{2} \text{V} + \underline{3} \text{Mn(ClO}_3)_2 \rightarrow$ $\underline{3} \text{Mn} + \underline{2} \text{V(ClO}_3)_3$	$\text{PbCl}_4 + \underline{2} \text{CuCO}_3 \rightarrow$ $\text{Pb(CO}_3)_2 + \underline{2} \text{CuCl}_2$														
<p>Balancing Chart ^{2 each (s)}</p> <table style="border: none;"> <tr> <td>V: 12</td> <td>V: 12</td> </tr> <tr> <td>Mn: 13</td> <td>Mn: 13</td> </tr> <tr> <td>ClO₃: 26</td> <td>ClO₃: 36</td> </tr> </table> <p style="text-align: center;">work = 2</p>	V: 1 2	V: 1 2	Mn: 1 3	Mn: 1 3	ClO ₃ : 2 6	ClO ₃ : 3 6	<p>Balancing Chart ^{2 each (s)}</p> <table style="border: none;"> <tr> <td>Pb: 1</td> <td>Pb: 1</td> </tr> <tr> <td>Cl: 4</td> <td>Cl: 24</td> </tr> <tr> <td>Cu: 12</td> <td>Cu: 12</td> </tr> <tr> <td>CO₃: 12</td> <td>CO₃: 2</td> </tr> </table> <p style="text-align: center;">work = 2</p>	Pb: 1	Pb: 1	Cl: 4	Cl: 2 4	Cu: 1 2	Cu: 1 2	CO ₃ : 1 2	CO ₃ : 2
V: 1 2	V: 1 2														
Mn: 1 3	Mn: 1 3														
ClO ₃ : 2 6	ClO ₃ : 3 6														
Pb: 1	Pb: 1														
Cl: 4	Cl: 2 4														
Cu: 1 2	Cu: 1 2														
CO ₃ : 1 2	CO ₃ : 2														

Name keyPeriod All

Quiz 20 - Polyatomic Ion Reaction Balancing

20 Points

Balance the following chemical reactions, showing balancing chart. No Work = No Credit!

Single Replacement Reaction	Double Replacement Reaction														
$\underline{3} \text{Mn} + \underline{2} \text{Cu(NO}_3)_2 \rightarrow$ $\underline{2} \text{Cu} + \underline{3} \text{Mn(NO}_3)_2$	$\text{TiBr}_4 + \underline{2} \text{PdSO}_4 \rightarrow$ $\text{Ti(SO}_4)_2 + \underline{2} \text{PdBr}_2$														
<p>Balancing Chart</p> <table style="border: none;"> <tr> <td>Mn: 13</td> <td>Mn: 13</td> </tr> <tr> <td>Cu: 12</td> <td>Cu: 12</td> </tr> <tr> <td>NO₃: 36</td> <td>NO₃: 26</td> </tr> </table> <p style="text-align: center;">work = 2</p>	Mn: 1 3	Mn: 1 3	Cu: 1 2	Cu: 1 2	NO ₃ : 3 6	NO ₃ : 2 6	<p>Balancing Chart</p> <table style="border: none;"> <tr> <td>Ti: 1</td> <td>Ti: 1</td> </tr> <tr> <td>Br: 4</td> <td>Br: 24</td> </tr> <tr> <td>Pd: 12</td> <td>Pd: 12</td> </tr> <tr> <td>SO₄: 12</td> <td>SO₄: 2</td> </tr> </table> <p style="text-align: center;">work = 2</p>	Ti: 1	Ti: 1	Br: 4	Br: 2 4	Pd: 1 2	Pd: 1 2	SO ₄ : 1 2	SO ₄ : 2
Mn: 1 3	Mn: 1 3														
Cu: 1 2	Cu: 1 2														
NO ₃ : 3 6	NO ₃ : 2 6														
Ti: 1	Ti: 1														
Br: 4	Br: 2 4														
Pd: 1 2	Pd: 1 2														
SO ₄ : 1 2	SO ₄ : 2														

Quiz 21 – Combination and Decomposition Reactions 20 Points

Determine the products and complete the following reactions. **No Work = No Credit!**

$\underline{3} \text{Zr}^{1+} + \underline{\quad} \text{N}_2 \rightarrow \underline{\quad} \text{Zr}_3\text{N}$ <p>A: $\underline{\text{Zr}} : + \underline{+1}$ B: $\underline{\text{N}} : - \underline{-3}$ <small>(2) each (2)</small></p>	<p style="text-align: center;"><i>AB Work</i></p> $\begin{array}{c} +1 \quad -3 \\ \text{Zr} \end{array} \text{N} \text{ (2)}$ $\text{Zr}_3\text{N} \text{ (2)}$	(8)
$\underline{\quad} \text{Cd}^{3+} + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad}$ <p>A: $\underline{\text{Cd}} : + \underline{+3}$ B: $\underline{\text{O}} : - \underline{-2}$ <small>(2) each (2)</small></p>	<p style="text-align: center;"><i>AB Work</i></p> $\begin{array}{c} +3 \quad -2 \\ \text{Cd} \end{array} \text{O}_2 \text{ (2)}$ $\text{Cd}_2\text{O}_3 \text{ (2)}$	(8)
$\underline{\quad} \text{MgI}_2 \rightarrow \underline{\quad} \text{Mg} + \underline{\quad} \text{I}_2$ <p style="text-align: center;"><small>(2) (2)</small></p>		(4)

Quiz 21 – Combination and Decomposition Reactions 20 Points

Determine the products and complete the following reactions. **No Work = No Credit!**

$\underline{3} \text{Mg} + \underline{\quad} \text{N}_2 \rightarrow \underline{\quad} \text{Mg}_3\text{N}_2$ <p style="text-align: center;"><small>A B</small></p> <p>A: $\underline{\text{Mg}} : + \underline{+2}$ B: $\underline{\text{N}} : - \underline{-3}$ <small>(2) each (2)</small></p>	<p style="text-align: center;"><i>AB Work</i></p> $\begin{array}{c} +2 \quad -3 \\ \text{Mg} \end{array} \text{N}_2 \text{ (2)}$ $\text{Mg}_3\text{N}_2 \text{ (2)}$	(8)
$\underline{\quad} \text{V}^{3+} + \underline{\quad} \text{Se} \rightarrow \underline{\quad} \text{V}_2\text{Se}_3$ <p>A: $\underline{\text{V}} : + \underline{+3}$ B: $\underline{\text{Se}} : - \underline{-2}$ <small>(2) each (2)</small></p>	<p style="text-align: center;"><i>AB Work</i></p> $\begin{array}{c} +3 \quad -2 \\ \text{V} \end{array} \text{Se} \text{ (2)}$ $\text{V}_2\text{Se}_3 \text{ (2)}$	(8)
$\underline{\quad} \text{TiCl}_4 \rightarrow \underline{\quad} \text{Ti} + \underline{2} \text{Cl}_2$ <p style="text-align: center;"><small>(2) (2)</small></p>		(4)

Name keyPeriod All**Quiz 21** – Combination and Decomposition Reactions

20 Points

Determine the products and complete the following reactions. **No Work = No Credit!**

$\underline{3} \text{ Fe}^{2+} + \underline{2} \text{ P} \rightarrow \underline{\quad} \text{ Fe}_3\text{P}_2$ <p>A: $\text{Fe} : + \frac{+2}{\textcircled{2}}$ B: $\text{P} : - \frac{-3}{\textcircled{2}}$ leach</p>	<p>AB Work</p> $\begin{array}{c} +2 \quad -3 \\ \text{Fe} \quad \text{P} \end{array} \textcircled{1}$ $\text{Fe}_3\text{P}_2 \textcircled{2}$	(8)
$\underline{2} \text{ Mn}^{3+} + \underline{3} \text{ S} \rightarrow \underline{\quad} \text{ Mn}_2\text{S}_3$ <p>A: $\text{Mn} : + \frac{+3}{\textcircled{2}}$ B: $\text{S} : - \frac{-2}{\textcircled{2}}$ leach</p>	<p>AB Work</p> $\begin{array}{c} +3 \quad -2 \\ \text{Mn} \quad \text{S} \end{array} \textcircled{1}$ $\text{Mn}_2\text{S}_3 \textcircled{2}$	
$\underline{\quad} \text{ NiBr}_2 \rightarrow \underline{\quad} \text{ Ni} + \underline{\quad} \text{ Br}_2$ <p style="text-align: center;">(2) 1 (2)</p>		(4)

Name keyPeriod All**Quiz 21** – Combination and Decomposition Reactions

20 Points

Determine the products and complete the following reactions. **No Work = No Credit!**

$\underline{3} \text{ Cr}^{2+} + \underline{2} \text{ P} \rightarrow \underline{\quad} \text{ Cr}_3\text{P}_2$ <p>A: $\text{Cr} : + \frac{+2}{\textcircled{2}}$ B: $\text{P} : - \frac{-3}{\textcircled{2}}$ leach</p>	<p>AB Work</p> $\begin{array}{c} +2 \quad -3 \\ \text{Cr} \quad \text{P} \end{array} \textcircled{1}$ $\text{Cr}_3\text{P}_2 \textcircled{2}$	(8)
$\underline{4} \text{ Cu}^{3+} + \underline{3} \text{ O}_2 \rightarrow \underline{2} \text{ Cu}_2\text{O}_3$ <p>A: $\text{Cu} : + \frac{+3}{\textcircled{2}}$ B: $\text{O} : - \frac{-2}{\textcircled{2}}$ leach</p>	<p>AB Work</p> $\begin{array}{c} +3 \quad -2 \\ \text{Cu} \quad \text{O} \end{array} \textcircled{1}$ $\text{Cu}_2\text{O}_3 \textcircled{2}$	(8)
$\underline{2} \text{ VBr}_3 \rightarrow \underline{2} \text{ V} + \underline{3} \text{ Br}_2$ <p style="text-align: center;">(2) (2)</p>		(4)

Name key Period 111
 Quiz 22 – Single Replacement Reactions 20 Points

Determine the products and complete the following reactions. No Work = No Credit!

$\underline{4} \text{ Al} + \underline{3} \text{ SnI}_4 \rightarrow \text{each (4)}$ $\underline{4} \text{ AlI}_3 + \underline{3} \text{ Sn}$ <p>A: <u>Al</u> : + <u>3</u> B: <u>Sn</u> each (5)</p> <p>C: <u>I</u> : - <u>1</u></p>	<p>AC Work (4) (10)</p> <p>+3 -1 Al I <u>AlI₃</u></p>
$\underline{3} \text{ Co}^{2+} + \underline{2} \text{ H}_3\text{PO}_4 \rightarrow \text{each (4)}$ $\underline{\text{Co}_3(\text{PO}_4)_2} + \underline{6} \text{ H}_2$ <p>A: <u>Co</u> : + <u>2</u> B: <u>H</u> each (5)</p> <p>C: <u>PO₄</u> : - <u>3</u></p>	<p>AC Work (4) (10)</p> <p>+2 -3 Co (PO₄) <u>Co₃(PO₄)₂</u></p>

Name key Period 111
 Quiz 22 – Single Replacement Reactions 20 Points

Determine the products and complete the following reactions. No Work = No Credit!

$\underline{4} \text{ Al} + \underline{3} \text{ SnCl}_4 \rightarrow \text{each (4)}$ $\underline{4} \text{ AlCl}_3 + \underline{3} \text{ Sn}$ <p>A: <u>Al</u> : + <u>3</u> B: <u>Sn</u> each (5)</p> <p>C: <u>Cl</u> : - <u>1</u></p>	<p>AC Work (4) (10)</p> <p>+3 -1 Al Cl <u>AlCl₃</u></p>
$\underline{3} \text{ Fe}^{2+} + \underline{2} \text{ H}_3\text{PO}_4 \rightarrow \text{each (4)}$ $\underline{\text{Fe}_3(\text{PO}_4)_2} + \underline{3} \text{ H}_2$ <p>A: <u>Fe</u> : + <u>2</u> B: <u>H</u> each (5)</p> <p>C: <u>PO₄</u> : - <u>3</u></p>	<p>AC Work (4) (10)</p> <p>+2 -3 Fe (PO₄) <u>Fe₃(PO₄)₂</u></p>

Determine the products and complete the following reactions. **No Work = No Credit!**

<p><u>4</u> 3 Al + <u>3</u> PbF₄ → <i>Teach (4)</i></p> <p style="text-align: center;"><u>4</u> AlF₃ + <u>3</u> 4 Pb</p> <p>A: <u>Al</u> : + <u>3</u> B: <u>Pb</u> <i>Teach (5)</i></p> <p>C: <u>F</u> : - <u>1</u></p>	<p style="text-align: center;"><i>AC Work</i> (1)</p> <p style="text-align: center;">+3 -1</p> <p style="text-align: center;">Al</p> <p style="text-align: center;"><u>AlF₃</u></p>
<p>_____ Cu²⁺ + _____ H₃PO₄ → <i>Teach (4)</i></p> <p style="text-align: center;">_____ <u>Cu₃(PO₄)₂</u> _____ <u>H₂</u></p> <p>A: <u>Cu</u> : + <u>2</u> B: <u>H</u> <i>Teach (5)</i></p> <p>C: <u>PO₄</u> : - <u>3</u></p>	<p style="text-align: center;"><i>AC Work</i> (7)</p> <p style="text-align: center;">+2 -3</p> <p style="text-align: center;">Cu(PO₄)</p> <p style="text-align: center;"><u>Cu₃(PO₄)₂</u></p>

(10)

(10)

Determine the products and complete the following reactions. **No Work = No Credit!**

<p><u>4</u> Al + <u>3</u> PbF₄ → <i>Teach (4)</i></p> <p style="text-align: center;"><u>4</u> AlF₃ + <u>3</u> Pb</p> <p>A: <u>Al</u> : + <u>3</u> B: <u>Pb</u> <i>Teach (5)</i></p> <p>C: <u>F</u> : - <u>1</u></p>	<p style="text-align: center;"><i>AC Work</i> (1)</p> <p style="text-align: center;">+3 -1</p> <p style="text-align: center;">Al</p> <p style="text-align: center;"><u>AlF₃</u></p>
<p><u>3</u> Cr²⁺ + <u>2</u> H₃PO₄ → <i>Teach (4)</i></p> <p style="text-align: center;">_____ <u>Cr₃(PO₄)₂</u> _____ <u>6</u> H₂</p> <p>A: <u>Cr</u> : + <u>2</u> B: <u>H</u> <i>Teach (5)</i></p> <p>C: <u>PO₄</u> : - <u>3</u></p>	<p style="text-align: center;"><i>AC Work</i> (1)</p> <p style="text-align: center;">+2 -3</p> <p style="text-align: center;">Cr(PO₄)</p> <p style="text-align: center;"><u>Cr₃(PO₄)₂</u></p>

(10)

(10)

Name keyPeriod All

Quiz 23 - Double Replacement Reactions

20 Points

Determine the products and complete the following reactions. No Work = No Credit!

$\underline{3} \text{PbI}_2 + \underline{\quad} \text{Al}_2\text{O}_3 \rightarrow$ $\underline{3} \text{PbO} + \underline{2} \text{AlI}_3$ <p>1/2 each (4)</p> <p>A: <u>Pb</u> : + <u>2</u> B: <u>I</u> : - <u>1</u></p> <p>C: <u>Al</u> : + <u>3</u> D: <u>O</u> : - <u>2</u></p>	<p>Cross Method Work</p> $\begin{array}{r} +2 \quad -2 \\ \text{Pb} \times \\ \text{Pb}_2\text{O}_2 \quad \text{PbO} \\ +3 \quad -1 \\ \text{Al} \times \\ \text{AlI}_3 \end{array}$ <p>Ⓛ</p>
$\underline{3} \text{Ca}(\text{NO}_3)_2 + \underline{2} \text{VPO}_4 \rightarrow$ $\underline{\quad} \text{Ca}_3(\text{PO}_4)_2 + \underline{2} \text{V}(\text{NO}_3)_3$ <p>1/2 each (4)</p> <p>A: <u>Ca</u> : + <u>2</u> B: <u>NO₃</u> : - <u>1</u></p> <p>C: <u>V</u> : + <u>3</u> D: <u>PO₄</u> : - <u>3</u></p>	<p>Cross Method Work</p> $\begin{array}{r} +2 \quad -3 \\ \text{Ca} \times \\ \text{Ca}_3(\text{PO}_4)_2 \\ +3 \quad -1 \\ \text{V} \times \\ \text{V}(\text{NO}_3)_3 \end{array}$ <p>Ⓛ</p>

Name keyPeriod All

Quiz 23 - Double Replacement Reactions

20 Points

Determine the products and complete the following reactions. No Work = No Credit!

$\underline{3} \text{SrF}_2 + \underline{\quad} \text{Cr}_2\text{O}_3 \rightarrow$ $\underline{3} \text{SrO} + \underline{2} \text{CrF}_3$ <p>1/2 each (4)</p> <p>A: <u>Sr</u> : + <u>2</u> B: <u>F</u> : - <u>1</u></p> <p>C: <u>Cr</u> : + <u>3</u> D: <u>O</u> : - <u>2</u></p>	<p>Cross Method Work</p> $\begin{array}{r} +2 \quad -2 \\ \text{Sr} \times \\ \text{Sr}_2\text{O}_2 \quad \text{SrO} \\ +3 \quad -1 \\ \text{Cr} \times \\ \text{CrF}_3 \end{array}$ <p>Ⓛ</p>
$\underline{3} \text{Mg}(\text{NO}_3)_2 + \underline{2} \text{FePO}_4 \rightarrow$ $\underline{\quad} \text{Mg}_3(\text{PO}_4)_2 + \underline{2} \text{Fe}(\text{NO}_3)_3$ <p>1/2 each (4)</p> <p>A: <u>Mg</u> : + <u>2</u> B: <u>NO₃</u> : - <u>1</u></p> <p>C: <u>Fe</u> : + <u>3</u> D: <u>PO₄</u> : - <u>3</u></p>	<p>Cross Method Work</p> $\begin{array}{r} +2 \quad -3 \\ \text{Mg} \times \\ \text{Mg}_3(\text{PO}_4)_2 \\ +3 \quad -1 \\ \text{Fe} \times \\ \text{Fe}(\text{NO}_3)_3 \end{array}$ <p>Ⓛ</p>

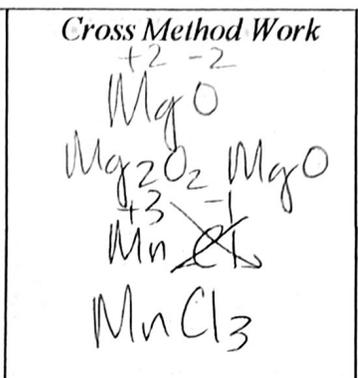
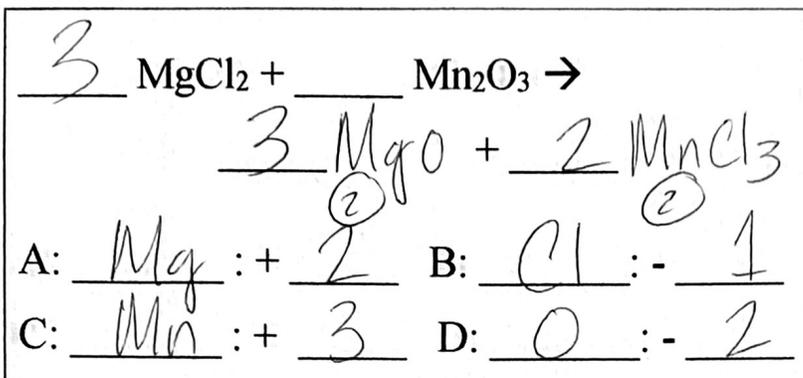
Name key

Period All

Quiz 23 - Double Replacement Reactions

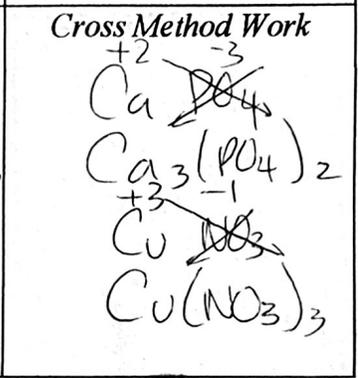
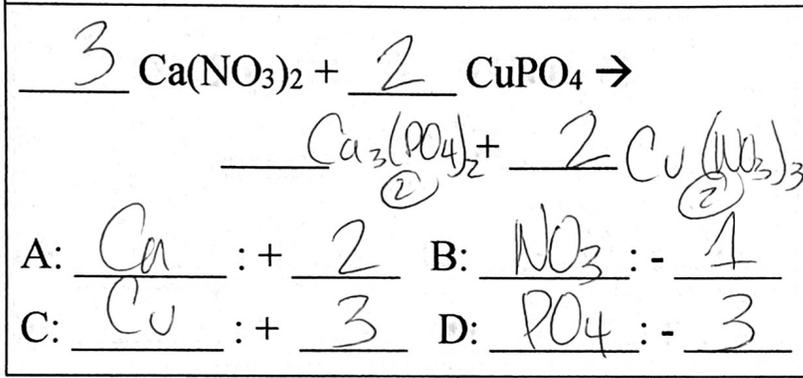
20 Points

Determine the products and complete the following reactions. No Work = No Credit!



1/2 each (4)

⓪
⓪



1/2 each (4)

⓪
⓪

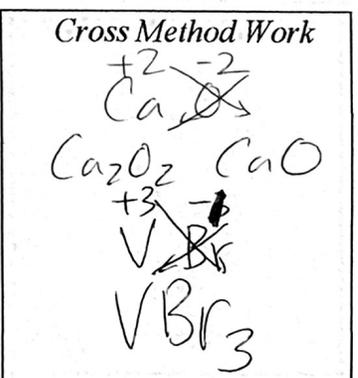
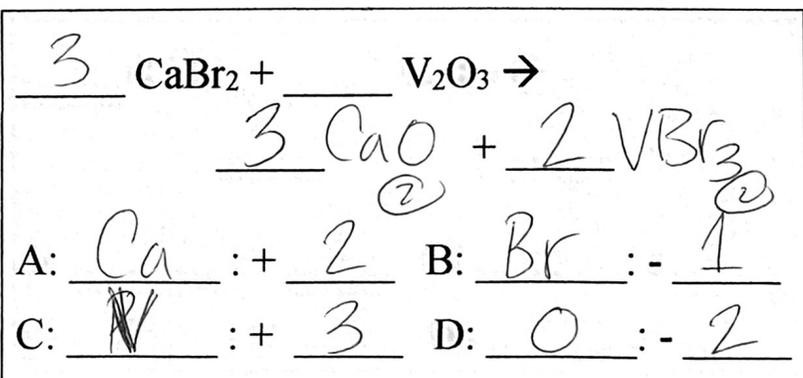
Name key

Period All

Quiz 23 - Double Replacement Reactions

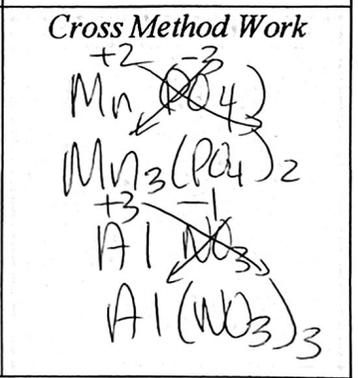
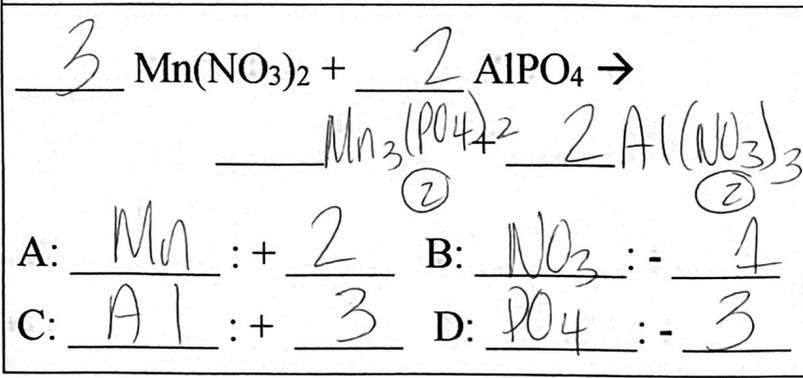
20 Points

Determine the products and complete the following reactions. No Work = No Credit!



1/2 each (4)

⓪
⓪



1/2 each (4)

⓪
⓪