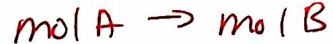
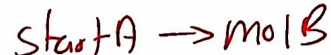


Overall Stoichiometry Chart

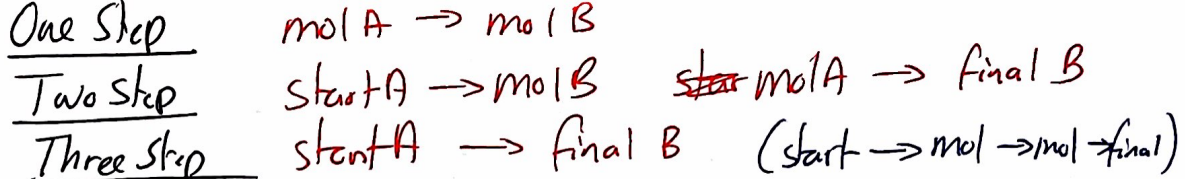
One Step



Two Step

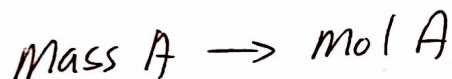


Three Step



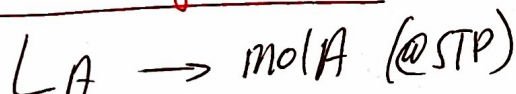
Step One

Mass (g)



<u>start</u> g A	1 mol A
	<u>Molar Mass</u> g A

Volume (gas @ STP)



<u>start</u> L A	1 mol A
	22.4 L A

Molarity (M)

$$\text{mol A} = \text{Molarity} \cdot V_{\text{soln}}$$

Step Two

Molar Ratio



<u>(step 1)</u> mol A	<u>Ratio</u> mol B
	<u>Ratio</u> mol A

Remember

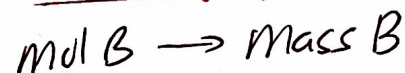
Mol Ratio requires labels
for mol A and mol B

Ex: 2.00 mol NaCl
(A)

Also round to 2 places after
decimal. 0.01, 230.10g

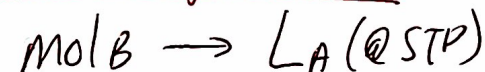
Step Three

Mass (g)



<u>(step 2)</u> mol B	<u>Molar Mass</u> g B
	1 mol B

Volume (gas @ STP)



<u>(step 2)</u> B	22.4 L B
	1 mol B

Molarity (M)

$$\text{Molarity} = \frac{(\text{step 2}) B}{V_{\text{soln}}}$$