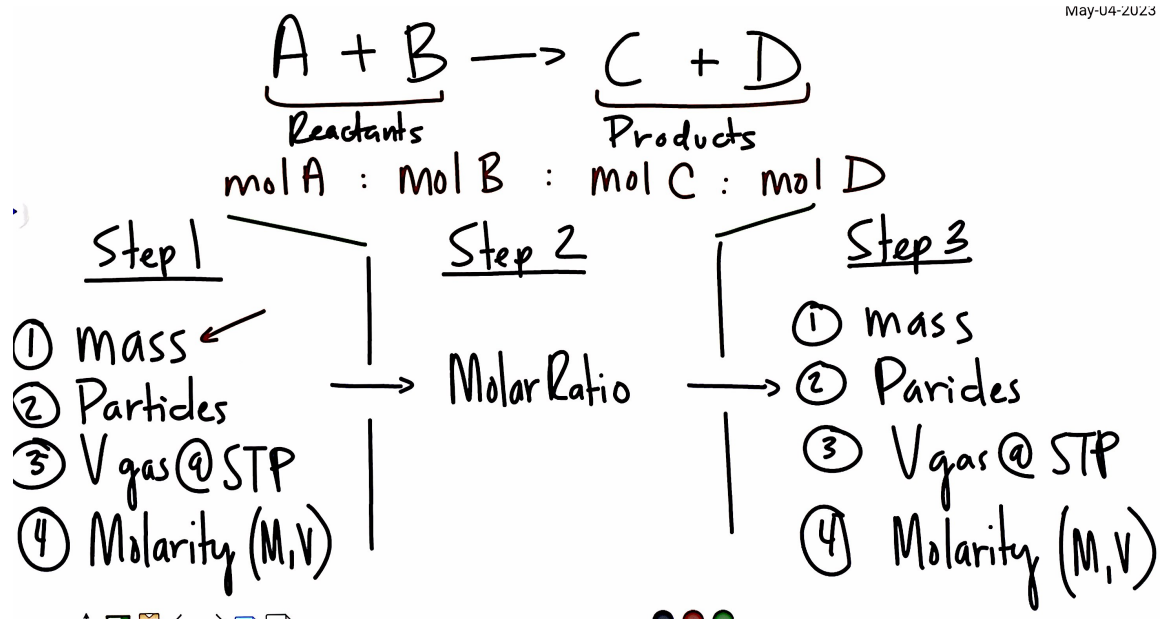
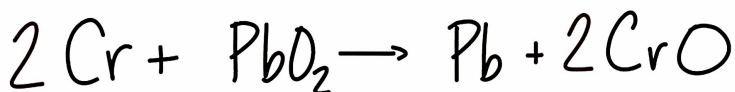
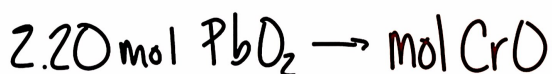




Step 2 : Molar Ratio

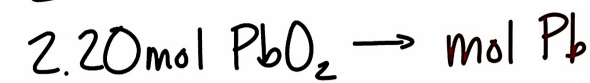


2 mol : 1 mol : 1 mol : 2 mol



$$\frac{2.20 \text{ mol PbO}_2}{1 \text{ mol PbO}_2} \times \frac{2 \text{ mol CrO}}{1 \text{ mol PbO}_2}$$

$$= 4.40 \text{ mol CrO}$$



$$\frac{2.20 \text{ mol PbO}_2}{1 \text{ mol PbO}_2} \times \frac{1 \text{ mol Pb}}{1 \text{ mol PbO}_2}$$

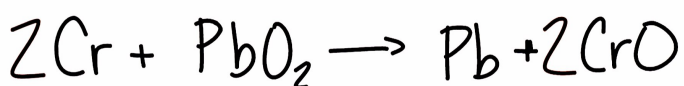
$$= 2.20 \text{ mol Pb}$$

Step 1 : get to mol (step 2)

Mass (1) $\rightarrow$ mol (2)

[Molar Mass, g/mol]

$$\frac{\text{mass} \rightarrow \text{mol}}{\text{start g} \quad \left| \quad \frac{1 \text{ mol}}{\text{MM g}}}\right.$$



given mass PbO_2 , want mol PbO_2

[Molar Mass $\text{PbO}_2 = 239.2 \text{ g/mol}$]

$$\begin{array}{l} 526.24 \text{ g PbO}_2 \rightarrow \\ \text{mol PbO}_2 \\ 526.24 \text{ g PbO}_2 \quad \left| \quad \frac{1 \text{ mol PbO}_2}{239.2 \text{ g PbO}_2} \right. \\ \hline = 2.20 \text{ mol PbO}_2 \end{array}$$