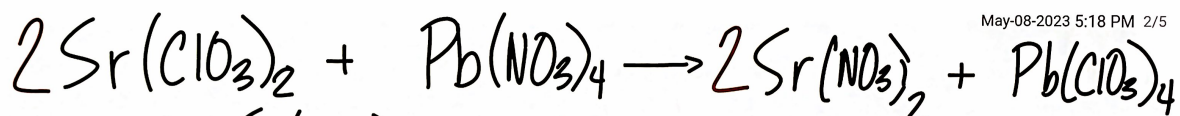


Mass Stoichiometry Chart

May-08-2023 5:10 PM 1/2

Step 1	Step 2	Step 3
$\text{Mass}_A \rightarrow \text{mol}_A$	$\text{mol}_A \rightarrow \text{mol}_B$	$\text{mol}_B \rightarrow \text{mass}_B$
$\frac{\text{start g A} \mid 1 \text{ mol A}}{\text{MM g A}}$	$\frac{\text{step 1 mol A} \mid * \text{mol B}}{* \text{mol A}}$	$\frac{\text{step 2 mol B} \mid \text{MM g B}}{1 \text{ mol B}}$
MM = Molar Mass (g/mol) (A)	* Molar Ratio (Bal. eqn.)	MM = Molar Mass (g/mol) (B)



May-08-2023 5:18 PM 2/5

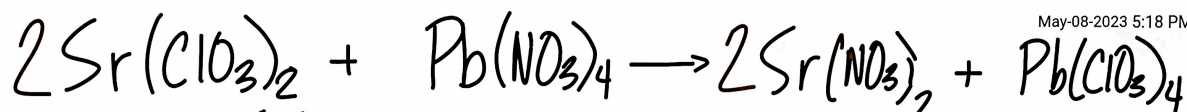
350.0g $\text{Sr}(\text{ClO}_3)_2$ convert to mass $\text{Pb}(\text{ClO}_3)_4$

(MM $\text{Sr}(\text{ClO}_3)_2 = 254.52 \text{ g/mol}$) (MM $\text{Pb}(\text{ClO}_3)_4 = 541.00 \text{ g/mol}$)

Step 1

Mass $\text{Sr}(\text{ClO}_3)_2 \rightarrow$
mol $\text{Sr}(\text{ClO}_3)_2$
(A)

$$\frac{350.0 \text{ g } \text{Sr}(\text{ClO}_3)_2 \mid 1 \text{ mol } \text{Sr}(\text{ClO}_3)_2}{254.52 \text{ g } \text{Sr}(\text{ClO}_3)_2 \mid \text{(MM)}} = 1.38 \text{ mol } \text{Sr}(\text{ClO}_3)_2$$



350.0g $\text{Sr}(\text{ClO}_3)_2$ convert to mass $\text{Pb}(\text{ClO}_3)_4$

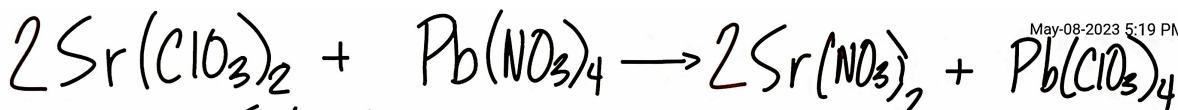
(MM $\text{Sr}(\text{ClO}_3)_2 = 254.52\text{g/mol}$) (MM $\text{Pb}(\text{ClO}_3)_4 = 541.00\text{g/mol}$)

Step 2
 $\text{mol Sr}(\text{ClO}_3)_2 \xrightarrow{\text{A}}$
 $\text{mol Pb}(\text{ClO}_3)_4 \text{ B}$

(step 1)

1.38 mol $\text{Sr}(\text{ClO}_3)_2$	1 mol $\text{Pb}(\text{ClO}_3)_4$
	2 mol $\text{Sr}(\text{ClO}_3)_2$

(Mol Ratios)
 $= 0.69 \text{ mol Pb}(\text{ClO}_3)_4$



350.0g $\text{Sr}(\text{ClO}_3)_2$ convert to mass $\text{Pb}(\text{ClO}_3)_4$

(MM $\text{Sr}(\text{ClO}_3)_2 = 254.52\text{g/mol}$) (MM $\text{Pb}(\text{ClO}_3)_4 = 541.00\text{g/mol}$)

Step 3
 $\text{mol Pb}(\text{ClO}_3)_4 \xrightarrow{\text{B}}$
 $\text{mass Pb}(\text{ClO}_3)_4 \text{ B}$

(step 2)

0.69 mol $\text{Pb}(\text{ClO}_3)_4$	541.00g $\text{Pb}(\text{ClO}_3)_4$
	1 mol $\text{Pb}(\text{ClO}_3)_4$

$= 373.29\text{g Pb}(\text{ClO}_3)_4$

