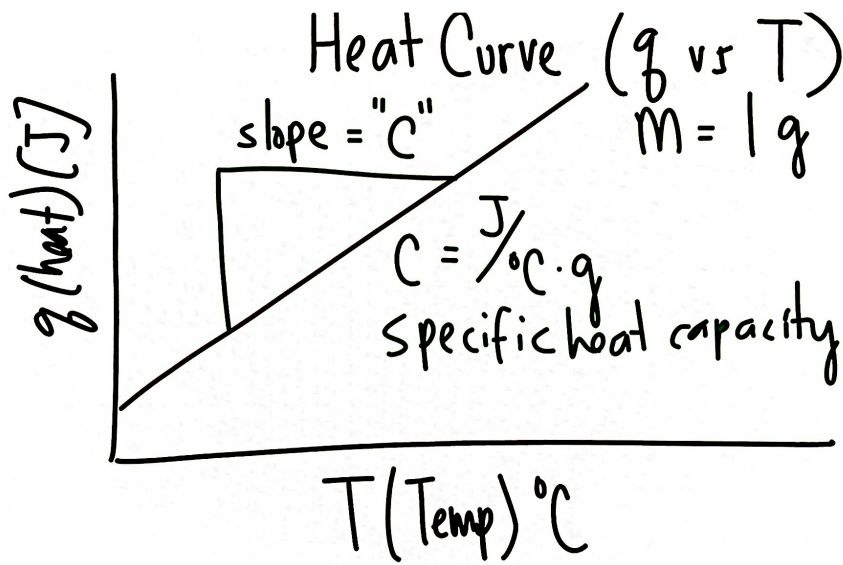
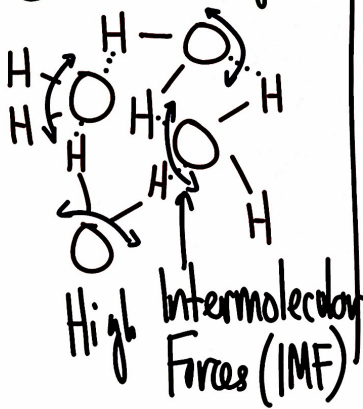


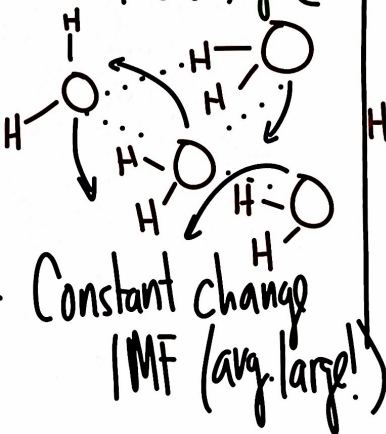


Heat Capacity
 q req for ΔT
 of 1°C
 [any m]

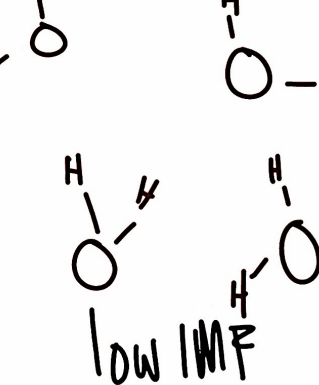
Solid
 $C = 2.09 \text{ J/g} \cdot ^{\circ}\text{C}$

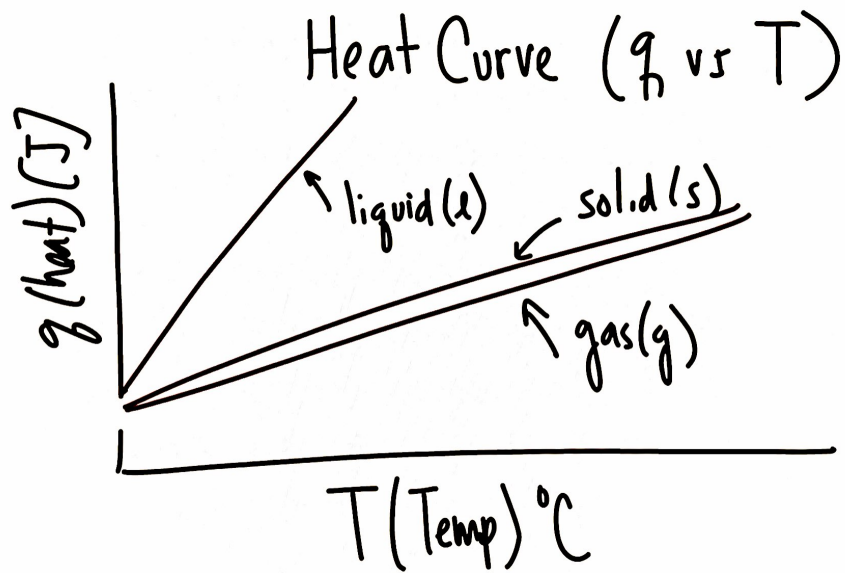


Liquid
 $C = 4.08 \text{ J/g} \cdot ^{\circ}\text{C}$



gas
 $C = 1.90 \text{ J/g} \cdot ^{\circ}\text{C}$





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Solid

$$C = 2.09 \text{ J/g} \cdot ^\circ\text{C}$$

$$m = 10 \text{ g}$$

$$\Delta T = 10^\circ\text{C}$$

$$q_s = C_s \cdot m \cdot \Delta T$$

$$= 2.09 \cdot 10 \cdot 10$$

$$q_s = 209 \text{ J}$$

liquid

$$C = 4.08 \text{ J/g} \cdot ^\circ\text{C}$$

$$m = 10 \text{ g}$$

$$\Delta T = 10^\circ\text{C}$$

$$q_l = C_l \cdot m \cdot \Delta T$$

$$= 4.08 \cdot 10 \cdot 10$$

$$q_l = 408 \text{ J}$$

gas

$$C = 1.90 \text{ J/g} \cdot ^\circ\text{C}$$

$$m = 10 \text{ g}$$

$$\Delta T = 10^\circ\text{C}$$

$$q_g = C_g \cdot m \cdot \Delta T$$

$$= 1.90 \cdot 10 \cdot 10$$

$$q_g = 190 \text{ J}$$