

Ionic Compound Nomenclature (naming)

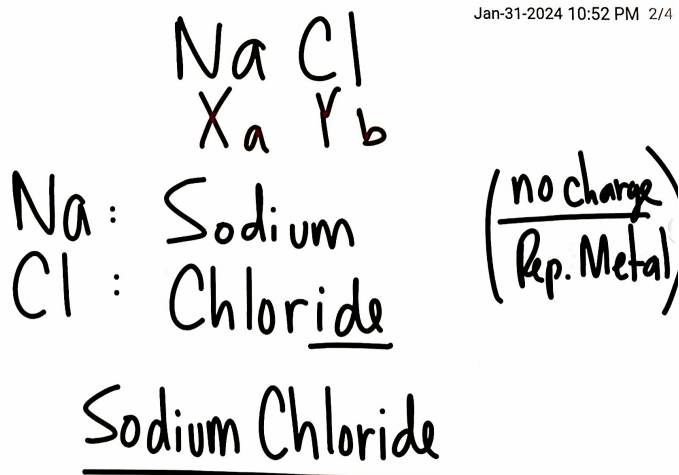
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$X_a Y_b$	
X: Rep. Metal (1A-3A)	<u>Metal</u> + <u>Non-Metal (-ide)</u>
Trans. Metal (1B-10B + 4A)	(X) (Y)
+ Poly. Ion	<u>Metal</u> + <u>Polyatomic</u>
Y: Rep. Anion	<u>Polyatomic</u> + <u>Non-Metal (-ide)</u>
- Poly. Ion	<u>Polyatomic</u> + <u>Polyatomic</u>

Naming Process

- ① Find X and Y
- ② Name X and Y
- ③ Find roman numeral
Trans. Metal
[Metal (numeral)]
- ④ Put Together Name

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Trans. Metals
(need charge)

Cu: +1, +2, or +3

Cu_2O_3

$$2x + -2 \cdot 3 = 0$$

$$2x = 6 \quad x = \underline{+3}$$

Cu_2O_3
 Cu^{3+} : Copper (III) (Trans. give +)
 +3 →

O: Oxide

Copper (III) Oxide

$\frac{\text{Ca}}{x} 3 \mid \frac{\text{N}}{y} 2$

Ca: Calcium Ion

N: Nitride Ion

Calcium | Nitride
 x y

Zn_2S Zn: +1

Zn^{1+} : Zinc (I) Ion

S: Sulfide Ion

Zinc (I) Sulfide