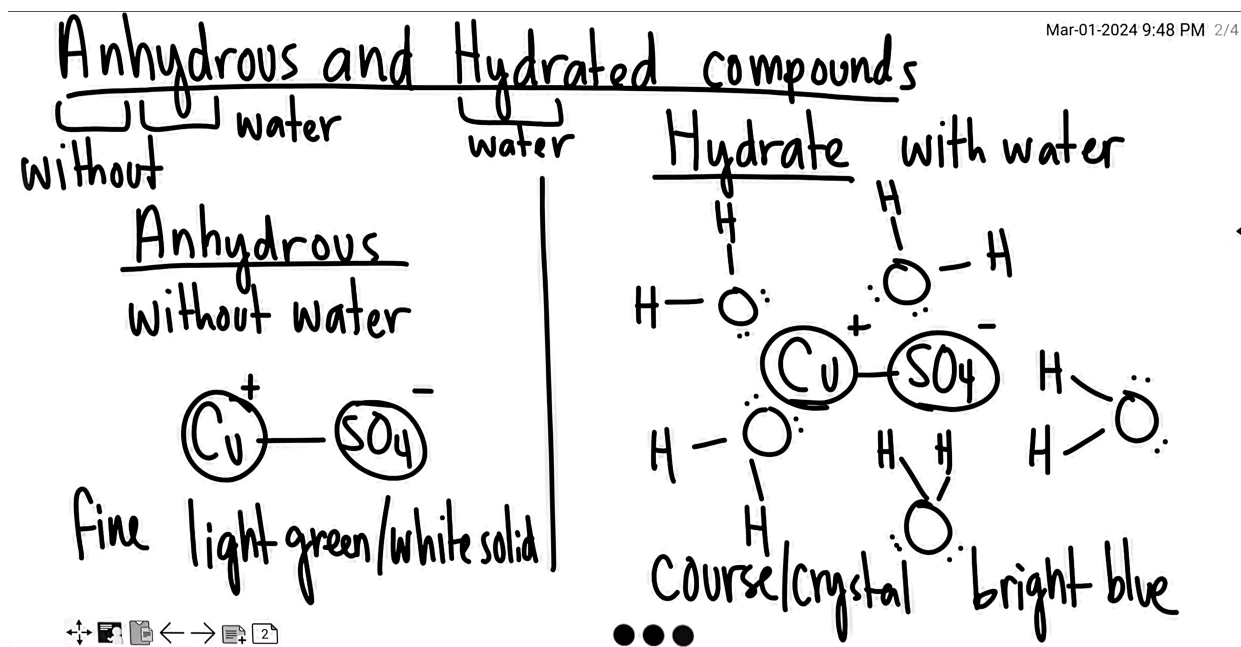
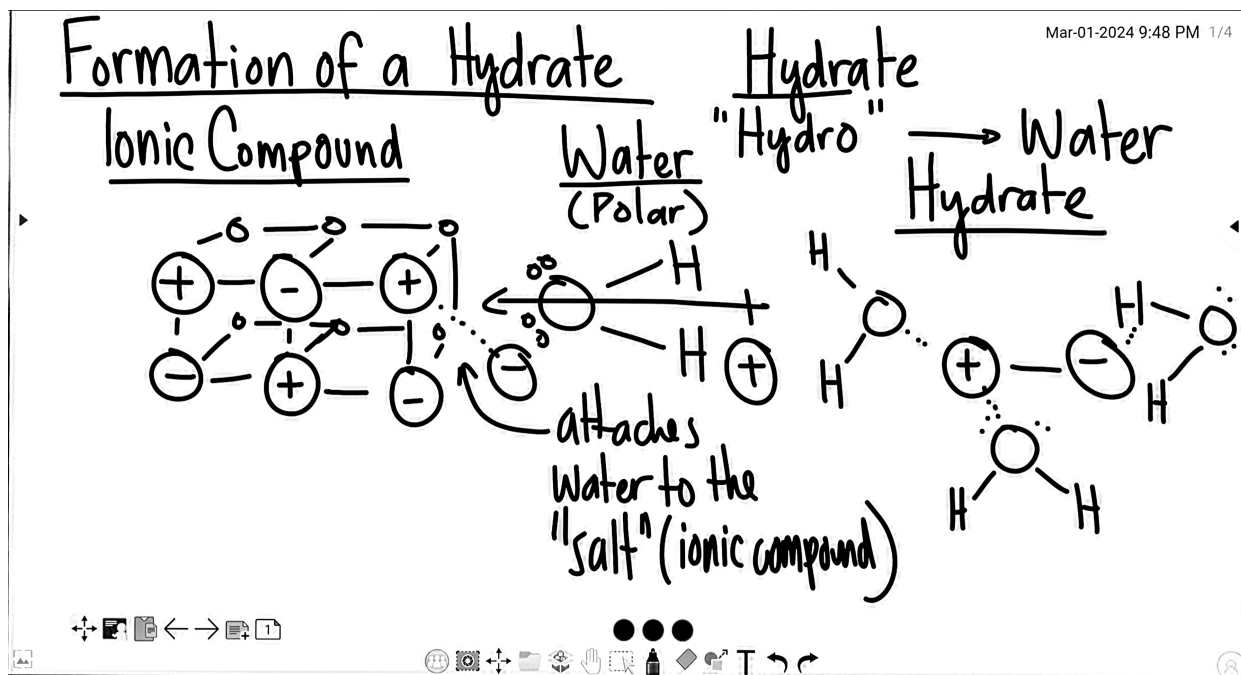




Anhydrous and Hydrated compounds

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<u>Molar/Atomic Ratio</u>	<u>Salt : Water (Ratio)</u>	
<u>Anhydrous Ratio</u>	<u>Hydrated Ratio</u>	
Salt:Water 1:0	1:1 <u>monohydrate</u> 1:4 <u>tetra</u>	
No (0) Water	(1)	1:5 <u>penta</u>
	1:2 <u>dihydrate</u>	(5)
	(2)	1:6 <u>hexa</u>
	1:3 <u>trihydrate</u>	(6)
	(3)	

← → 3

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Hydrate Examples

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$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
Ratio: 1 | 5
 CuSO_4 : Copper(II) Sulfate
 $5\text{H}_2\text{O}$: Penta hydrate
 Copper(II) Sulfate Penta hydrate

Fe_2O_3 Iron(III) Oxide
 MnCl_2 Manganese(II) Chloride
 Na_2SO_4 Sodium Sulfate
 $\text{Zn}_2\text{S}_2\text{O}_3$ Zinc(I) thiosulfate
 $\text{Bi}(\text{NO}_2)_3$ Bismuth(III) Nitrite
 $\text{Al}_2(\text{SO}_3)_3$ Aluminum Sulfite

← → 4

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