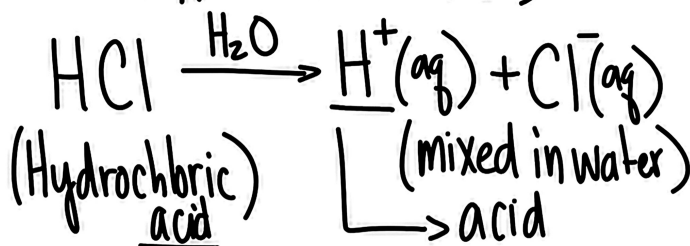


Acids and Bases

Acid (HX or HXO_2)

- Molecule that donates (loses) H^+ in water (H_2O)



Types of Acid

Binary (HX) Acid



Binary = 2 atoms

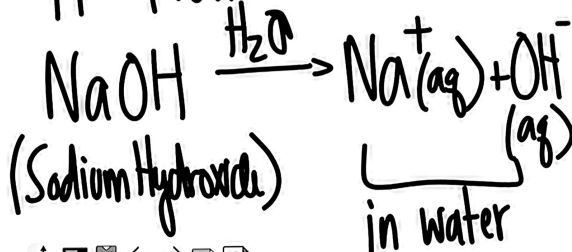
Poly (HXO_2) Acid
polyatomic ion



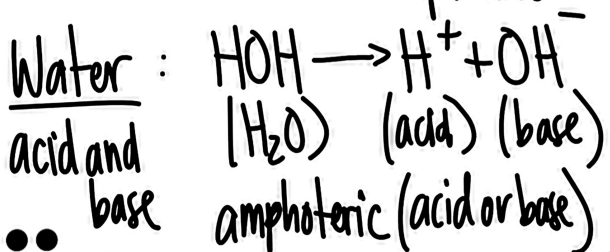
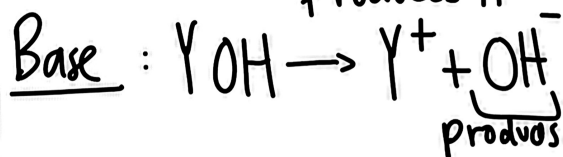
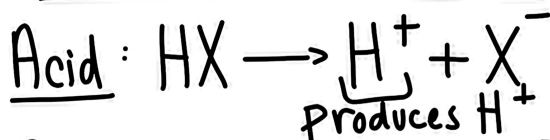
Acids and Bases

Bases (YOH)

- Compound containing OH^{1-} (hydroxide) or accepts a H^+ from an acid.



Compare Acids and Bases

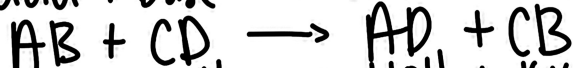


Acid Base Neutralization

Reacting an acid with a base
to produce (make) a salt and water



general: acid + base $\longrightarrow$ salt + water



Neutralize: Reduce

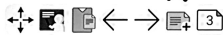
H^+ ions in solution
(mixture)

(add base to an acid)

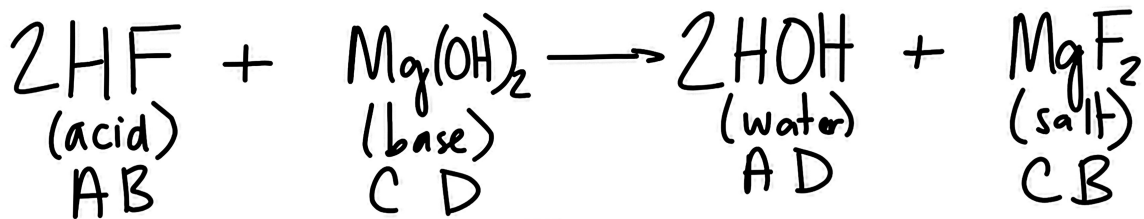
salt*

Ionic compound
formed from
an acid/base
neutralization

(Double
Replace)



Acid Base Neutralization



A: H
D: OH

acid
base

C: Mg
B: F

salt

Ratio: 2 : 1 : 2 : 1

Acid Base Neutralization[^]

