

Assign. 6A

Selected Answers

Nitrogen - 14

Mass # = 14 ($p^+ + n^0$)

Atomic Mass = 14 amu

Atomic Mass (g)

$$1 \text{ amu} = 1.67 \times 10^{-24} \text{ g/amu}$$

$$\text{Atomic Mass}_{(\text{amu})} \times 1.67 \times 10^{-24} \text{ g/amu}$$

$$14 \times 1.67 \times 10^{-24} \text{ g} = 2.338 \times 10^{-23} \text{ g}$$

Molar Mass = Atomic Mass in g/mol

$$\text{Atomic Mass} = \frac{14 \text{ amu}}{\text{atom}} \quad \text{Molar Mass} = 14 \text{ g/mol}$$