60 Review

Isotopes : Manganese-53 ← Mass # ($p^+ + n^0$)

Atomic mass (amu) — 53 amu (1 atom = 53 amu)

Molar Mass (g/mol) — 53 g/mol (53 g = 1 mol)

Avg. Molar Mass — 54.94 g/mol (54.94 g = 1 mol) ← Periodic Table

60 Review

Phosphoric Acid $[H_3PO_4]$ $(PO_4)_1 = 1$

$\rightarrow 1$ $(PO_4)_2 = 2$

Hydrogen [H]	3	1.01 g
Phosphate [PO_4]	1	94.97 g
Phosphoric Acid [H_3PO_4]		98.00 g/mol

PO_4	
P : 1	30.97 g
O : 4	16.00 g
PO_4 = 94.97 g	

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6D Review

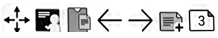
Convert $59.2\,L\ SO_2(g) \rightarrow mol\ SO_2(g)$

[1 mol = 22.4 L]

$$59.2\,\cancel{L\ SO_2} \times \frac{1\ mol\ SO_2}{22.4\,\cancel{L\ SO_2}} = 2.64\ mol\ SO_2(g)$$

Vol → mol

$$(start)\ L_A \times \frac{1\ mol_A}{22.4\,L_A} = \_\_\_\_\ mol_A$$



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$$\underline{1}\ H_2CO_3 + \underline{2}\ NaOH \rightarrow \underline{2}\ HOH + \underline{1}\ Na_2CO_3$$

1 mol : 2 mol : 2 mol : 1 mol

Convert 59.3 L NaOH to mol NaOH | mol NaOH → mol HOH | mol HOH → Vol (L) HOH

$$59.3\,\cancel{L\ NaOH} \times \frac{1\ mol\ NaOH}{22.4\,\cancel{L\ NaOH}} \times \frac{2.65\,\cancel{mol\ NaOH}}{1} \times \frac{2\ mol\ HOH}{2\,\cancel{mol\ NaOH}} \times \frac{2.65\,\cancel{mol\ HOH}}{1} \times \frac{22.4\ L\ HOH}{1\,\cancel{mol\ HOH}} = 59.3\ L\ HOH(g)$$