

College Prep Chemistry of the Earth System

Assignment 1F – Isotope Series and Atomic Mass

20 Point

Complete the following Chart based on the isotope Series No Work = No Credit

Avg. Atomic
Mass

Isotope Notation	Atomic #	Mass #	Protons (p ⁺)	Electrons (e ⁻)	Neutrons (n ⁰)	Mass #	
Nickel-59							
Argon-40							
9 4	Be						
107 47	Ag						
Gallium-69							
Gallium-71							
35 17	Cl	17	35	17	17	35-17 = 18	35 amu
37 17		17	37	17	17	37-17 = 20	37 amu
Potassium-39						Atomic Mass =	
Potassium-40						Mass # In amu	
Potassium-41							
204 82	Pb						
206 82		Pb					
207 82	Pb						
208 82		Pb					

$n^0 = \text{Mass \#} - \text{Atomic \#}$

Atomic # Mass # # p⁺ # e⁻ # n⁰ Atomic Mass (amu)