

Name _____ Period _____

Freshmen Transition

Assignment 1H – Atomic Mass

20 Points

Answer the following questions based on the in class notes

What are the two particles that determine the mass on an atom (<i>in g or amu</i>)?
Do electrons have mass, and if it does, why is the mass not counted in the atomic mass?
What is the difference between mass number and atomic mass?
Why is the mass of an atom given in amu (<i>atomic mass units</i>) over grams (<i>g</i>)?

Complete the following chart based on the atomic structure and atomic mass

Atomic Isotope	Atomic Number	# p ⁺	p ⁺ Mass (amu) *	# n ⁰ *	n ⁰ Mass (amu) *	Mass Number	Atomic Mass (amu) *
Magnesium-24	12	12	12amu	24-12 12	12amu	24	24amu
Vanadium-51			*# p ⁺ (amu)		*Mass# - Atomic #		
Bromine-80			*# n ⁰ (amu)		*Mass# (amu)		
Tin-119							
Tungsten-184							
Radium-226							

Atomic #

24
= Mg

12

Atomic #

(PT)

12
Mg