Isotopes: same # p^+, different # n^0 ($p^+ = e^-$)

[n^0 stabilize atom]

$^{35}_{17}Cl$ ← Mass #, Atomic #

Chlorine-35

Mass # = $p^+ + n^0$

p^+: 17 e^-: 17

n^0 = Mass # − Atomic # = 35 − 17 = 18

$^{37}_{17}Cl$ ← Mass #, Atomic #

Chlorine-37

Change

p^+: 17 e^-: 17

n^0 = 37 − 17 = 20

Isotope (atomic) mass

Mass of atom in atomic mass unit (amu)

1 amu = Mass of 1 p^+ or 1 n^0

$^{35}_{17}Cl$ ← Mass # (# p^+ + # n^0)

Atomic Mass

Mass # in amu (Mass of $p^+ + n^0$)

$^{35}_{17}Cl$

Mass # = 35 (count)

Atomic Mass = 35 amu (measurement)

$^{37}_{17}Cl$ Mass # 37 Atomic Mass = 37 amu