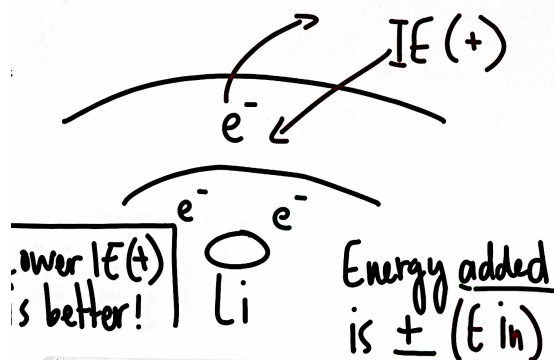


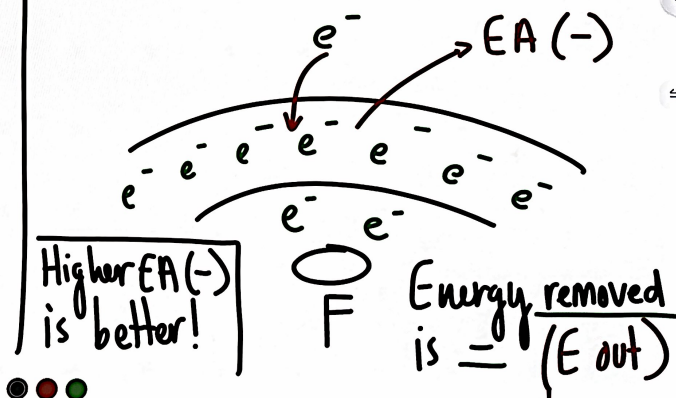
Ionization Energy

Energy required to remove electron from neutral atom



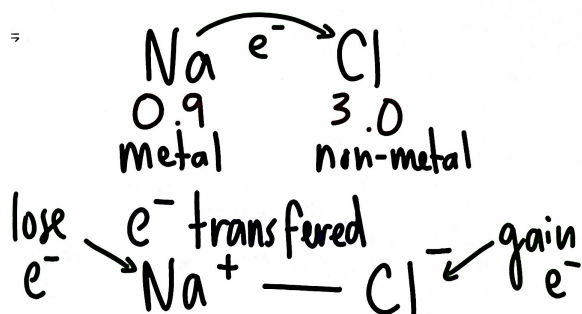
Electron Affinity

Energy released when electron added to neutral atom



Electronegativity

An atom's "ability" to attract electrons towards itself in a chemical bond (two atom interaction)



Selected Electroneg. Values

Metals

Li	1.0
K	0.8
Mg	1.2
Ca	1.0
Al	1.5
(lose e^-)	

Non-Metals

N	3.0
O	3.5
S	2.5
F	4.0
Cl	3.0
(gain e^-)	

Trend Summary

Ionization Energy

Increases across table

Decrease down table

Electron Affinity

Increase across table

Decrease down table

Atom Size

Decrease (a little) across table

Increase (a lot) down table

Ion Size

Cations(+) smaller | Anions(-)
larger

Electronegativity

Increase across table

Decrease down group