

Name _____ Period _____

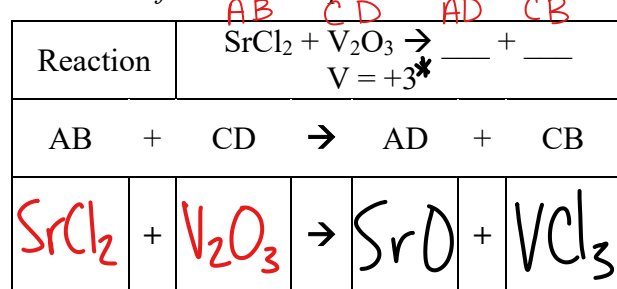
College Prep Chemistry of the Earth

Assignment 4K – Writing Double Replacement Reactions

20 Points

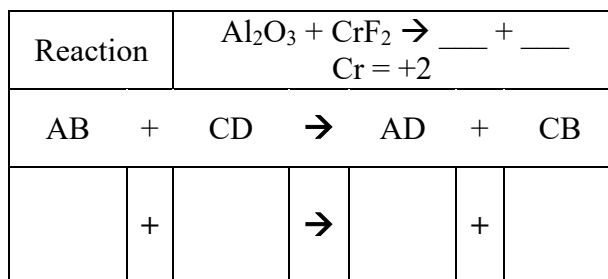
For the following reactions, write the products of each reaction following the template provided.

Show work for ionic compound



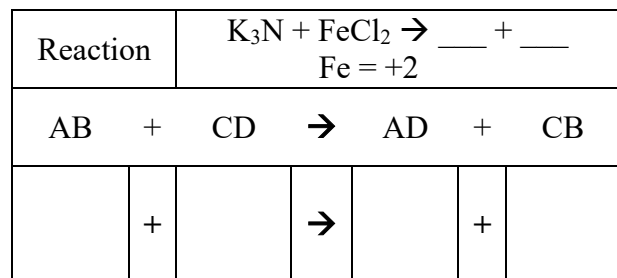
	A	D	C	B
Element or Ion	Sr	O	V	Cl

Ion Charge	+2	-2	+3*	-1
Cross Method	Sr O		V Cl	
Final Formula	Sr ₂ O ₂		VCl ₃	
Redu.	> SrO			



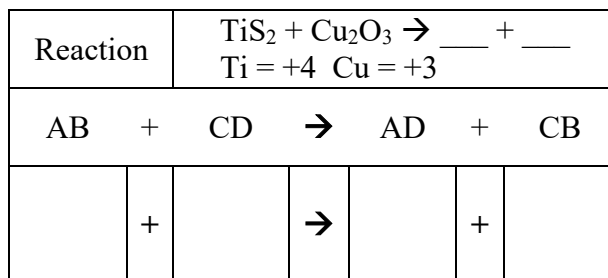
	A	D	C	B
Element or Ion				

Ion Charge				
Cross Method				
Final Formula				



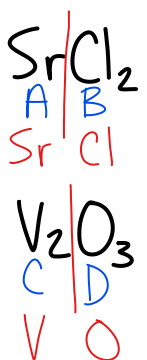
	A	D	C	B
Element or Ion				

Ion Charge				
Cross Method				
Final Formula				



	A	D	C	B
Element or Ion				

Ion Charge				
Cross Method				
Final Formula				



Min. Work!
Show answer!

Reaction	$\text{AlP} + \text{CuO} \rightarrow \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$ $\text{Cu} = +2$					
AB	+	CD	$\rightarrow$	AD	+	CB
	+		$\rightarrow$		+	

	A	D	C	B
Element or Ion				

Ion Charge				
Cross Method				
Final Formula				

Reaction	$\text{WN}_2 + \text{ZnF} \rightarrow \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$ $\text{W} = +6 \quad \text{Zn} = +1$					
AB	+	CD	$\rightarrow$	AD	+	CB
	+		$\rightarrow$		+	

	A	D	C	B
Element or Ion				

Ion Charge				
Cross Method				
Final Formula				

Reaction	$\text{SnSe}_2 + \text{AgI} \rightarrow \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$ $\text{Sn} = +4 \quad \text{Ag} = +1$					
AB	+	CD	$\rightarrow$	AD	+	CB
	+		$\rightarrow$		+	

	A	D	C	B
Element or Ion	Sn	I	Ag	Se

Ion Charge	+4	-1	+1	-2
Cross Method	Sn I SnI_4		Ag Se Ag_2Se	
Final Formula				

Reaction	$\text{Pb}_3\text{N}_4 + \text{CuBr}_2 \rightarrow \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$ $\text{Pb} = +4 \quad \text{Cu} = +2$					
AB	+	CD	$\rightarrow$	AD	+	CB
	+		$\rightarrow$		+	

	A	D	C	B
Element or Ion				

Ion Charge				
Cross Method				
Final Formula				

SnSe_2
 AgI
 SnI_4
 Ag_2Se
 min work!
 Write Answer!