Name Selected Answers Period _____

College Prep Chemistry of the Earth

Assignment 5A – Mass of an Atom 20 Points

Answer the following questions based on the in class presentation

Define Mass Number	Define Atomic Mass Unit
Define Atomic Mass	**Define Average Atomic Mass**

Subatomic Particle Data Tables

Mass Proton (p^+)	Mass Neutron (n^o)	Atomic Mass Unit	Mass Number
$1.67x10^{-24}g/p^+$	$1.67x10^{-24}g/e^-$	$1.67x10^{-24}g/amu$	$\#p^+ + \#n^o$

Complete the following chart based on the isotopes shown below [1amu = $1.67x10^{-24}$g]

Isotope	Mass Number (amu)	X	Atomic Mass Unit (g/amu)	=	Atomic Mass (g)
Nitrogen-14 (Mass #)	14 amu	X	$1.67x10^{-24}$ g/amu	=	$2.338x10^{-23}$ g
Magnesium-24		X		=	
Titanium-48		X		=	
Selenium-80		X		=	

Periodic Table

7
N
14.01

(Atomic Mass (amu))

Isotope	Average Atomic Mass (amu)	X	Atomic Mass Unit (g/amu)	=	Atomic Mass (g)
Nitrogen	14.01 amu	X	$1.67x10^{-24}$ g/amu	=	$2.340x10^{-23}$ g
Magnesium		X		=	
Titanium		X		=	
Selenium		X		=	

On Calculator x10 = EE or Exp or $x10^x$ Ex: 1.67 EE −24 =