

Name _____ Period _____

College Prep Chemistry of the Earth

Assignment 5B – The Mol - Atomic and Molar Mass

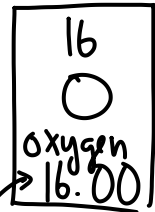
20 Points

Answer the following questions based on the in class discussion

Define Mol	Define Avogadro's Number

Determine the atomic mass (amu), and molar mass (g/mol) for the following isotopes, and the average atomic mass (amu and g/mol) based on the periodic table.

Isotope	Oxygen-16	Oxygen-17	Oxygen-18
Atomic Mass (amu)	16 amu	17 amu	18 amu
Molar Mass (g/mol) *	16 g/mol	17 g/mol	18 g/mol
Avg Atomic Mass (g/mol, periodic table) **	16.00 g/mol	← Use # on PT (2 pl. after dec.)	


 Avg. Atomic Mass
 g/mol

Mass # →
 Oxygen-16
 * Molar Mass
 $\frac{16 \text{ g}}{1 \text{ mol}}$

Isotope	Magnesium-24	Magnesium-25	Magnesium-26
Atomic Mass (amu)			
Molar Mass (g/mol)			
Avg Atomic Mass (g/mol, periodic table)			

Isotope	Iron-54	Iron-56	Iron-58
Atomic Mass (amu)			
Molar Mass (g/mol)			
Avg Atomic Mass (g/mol, periodic table)			

Isotope	Antimony-121	Antimony-123	Antimony-125
Atomic Mass (amu)			
Molar Mass (g/mol)			
Avg Atomic Mass (g/mol, periodic table)			