

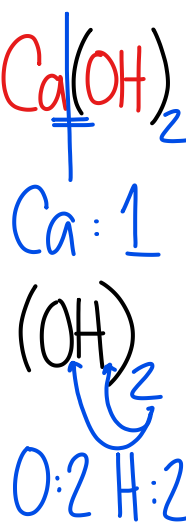
Name _____ Period _____

College Prep Chemistry of the Earth

Assignment 5E – Molar Mass (*multiple polyatomic ions*)

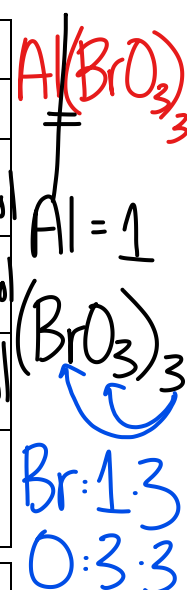
20 Points

For the following ionic compounds or covalent molecules, calculate the molar mass



Calcium Hydroxide $[\text{Ca(OH)}_2]$		
Element	#	Molar Mass
Calcium (Ca)	1	40.08 g/mol
Oxygen (O)	2	16.00 g/mol
Hydrogen (H)	2	1.01 g/mol
Calcium Hydroxide $[\text{Ca(OH)}_2]$		$= 74.10 \text{ g/mol}$

Aluminum Bromate $[\text{Al(BrO}_3)_3]$		
Element	#	Molar Mass
Aluminum (Al)	1	26.98 g/mol
Bromine (Br)	3	79.90 g/mol
Oxygen (O)	9	16.00 g/mol
Aluminum Bromate $[\text{Al(BrO}_3)_3]$		$= 410.68 \text{ g/mol}$



Iron(III)Chromate $[\text{Fe}_2(\text{CrO}_4)_3]$		
Element	#	Molar Mass
Iron (Fe)		
Chromium (Cr)		
Oxygen (O)		
Iron(III)Chromate $[\text{Fe}_2(\text{CrO}_4)_3]$		

Beryllium Phosphate $[\text{Be}_3(\text{PO}_4)_2]$		
Element	#	Molar Mass
Beryllium (Be)		
Phosphorous (P)		
Oxygen (O)		
Beryllium Phosphate $[\text{Be}_3(\text{PO}_4)_2]$		

Lead(IV)Thiosulfate $[\text{Pb(S}_2\text{O}_3)_2]$		
Element	#	Molar Mass
Lead (Pb)		
Sulfur (S)		
Oxygen (O)		
Lead(IV)Thiosulfate $[\text{Pb(S}_2\text{O}_3)_2]$		

Tungsten(IV)Cyanide $[\text{W(CN)}_6]$		
Element	#	Molar Mass
Tungsten (W)		
Carbon (C)		
Nitrogen (N)		
Tungsten(IV)Cyanide $[\text{W(CN)}_6]$		