

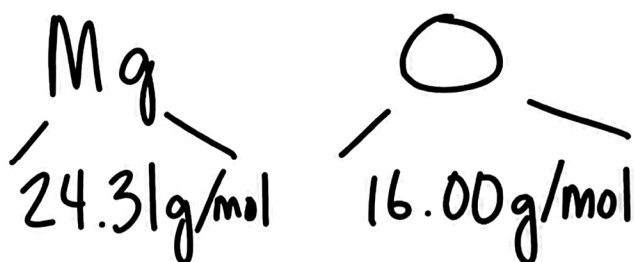
Molar Mass

Single Element

use # for mass on periodic table

12	← Atomic #
Mg	← Symbol
Magnesium	← Name
24.31	← Molar Mass

Molar Mass is additive



$$\begin{aligned} \text{MgO} &= 24.31 \text{ g/mol} + 16.00 \text{ g/mol} \\ &= \underline{40.31 \text{ g/mol}} \end{aligned}$$

Molar Mass w/ subscripts

Binary $A_x B_y$

Molar Mass Eqn

$$\text{MM} = x \cdot A + y \cdot B$$

A and B are the Molar Mass
x and y are the subscripts

Ex: AlCl_3 (Al_1Cl_3)

$$\text{Al}: 26.98 \text{ g/mol} \quad \text{Cl}: 35.45 \text{ g/mol}$$

$$\begin{aligned} \text{MM}_{\text{AlCl}_3} &= \underbrace{1 \cdot 26.98 \text{ g/mol}}_{\text{Al}} + \underbrace{3 \cdot 35.45 \text{ g/mol}}_{\text{Cl}} \\ &= \underline{133.33 \text{ g/mol}} \end{aligned}$$