Name ______ Period _____

Lab 8 – Stoichiometry of an Ideal Gas

50 Points

Introduction

In this lab you will be performing a gas producing reaction between sodium bicarbonate $[NaHCO_3(s)]$ and acetic acid $[HC_2H_3O_2(aq)]$ in an aqueous solution producing CO_2 gas

$$NaHCO_3(s) + HC_2H_3O_2(aq) \rightarrow NaC_2H_3O_2(aq) + CO_2(g) + H_2O(l)$$

By using stoichiometry you will calculate the amount of sodium bicarbonate needed to completely react with 100mL of a 1.00M $HC_2H_3O_2(aq)$ solution (0.100mol $HC_2H_3O_2(aq)$). The product of the lab, carbon dioxide gas, will be collected in a balloon, and measured using a flexible ruler. The volume of a balloon is then calculated with the circumference to find the total volume of $CO_2(g)$.

The experimental volume of $CO_2(g)$ is then compared to the theoretical volume based on stoichiometry to determine the accuracy and precision of the lab.

Procedure

- 1. Weigh the mass of an empty balloon using the laboratory balance. Add NaHCO₃(s) to the balloon, subtracting the mass of the balloon until you obtain the mass calculated in pre-lab question 1;
- 2. Measure out 100.0mL of 1.00mol/L HC₂H₃O₂(aq) and add to a 125mL Erlenmeyer flask. Fill the flask to the top with deionized water, and allow the acid and water to mix;
- 3. Place balloon lip over the lip of the flask while not allowing the sodium bicarbonate to fall into the balloon;
- 4. Hold the lip of the balloon and allow the acid to flow into the balloon to wash out the sodium bicarbonate. Allow the acid to flow back into the beaker, completely draining the acid from the balloon. Allow the reaction to proceed while holding the balloon;
- 5. Once the reaction completes, carefully remove the balloon lip from the flask, and tie a knot in the balloon. Measure the circumference of the balloon, and record on data table;
- 6. Clean out all laboratory glassware and return supplies to the stock table. Carefully put a hole in the balloon, and throw the balloon in the large trash can.

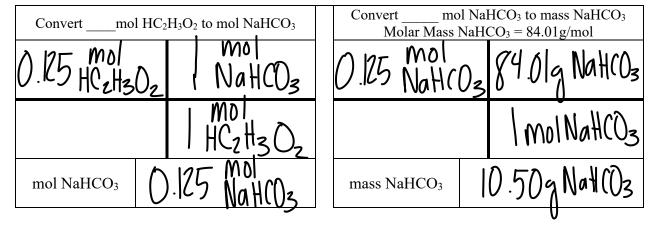
Pre-Lab Calculations

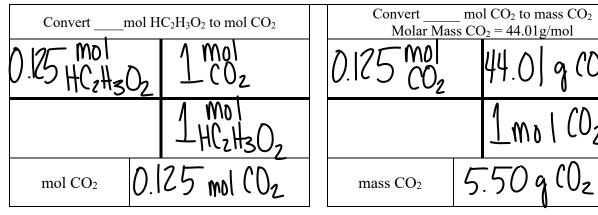
Balanced Reaction: NaHCO₃(s) + HC₂H₃O₂(aq) \rightarrow NaC₂H₃O₂(aq) + CO₂(g) + H₂O(l)

Calculate the mol HC₂H₃O₂ (acetic acid, vinegar)

Starting Values (Measurements)	Molarity Equation	Solve for mol HC ₂ H ₃ O ₂	Mol $HC_2H_3O_2$
Molarity = 1.00mol/L	Molarity = mol / L	$\underset{HC_2H_3O_2}{\text{mol}} _{100} \frac{\text{Mol}}{l} \times 0. 25l$	0.125mal
$V_{\text{solution}} = 0.100L$ $(100mL)$	$mol = Molarity \cdot L$	$\underset{HC_2H_3O_2}{\text{mol}} \underbrace{\uparrow})) \underbrace{\downarrow} \times (). 25 ($	mol HC ₂ H ₃ O ₂

Calculate Mass (g) of NaHCO₃ to perform reaction

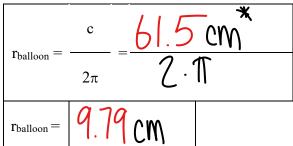




Lab Data Set

Laboratory	$mol\ CO_2(g)$	Pressure Room	Temperature	Ideal Gas
Measurement	(Calc Above)	(P_{room})	Room (T_{room})	Constant (R)
	0.125 mol	1.07 alm	2%.2K	0.0821 L atm mol K
Laboratory	Circumference	Volume of Flask	Radius of Balloon	Volume of Balloon
Measurement	of Balloon (cm)	(mL)	(cm, calculated)	(cm, calculated)
Experimental Data Value	61.5 cm*	125 mL	9.79 cm	3930.40 3.931
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Post-Lab Calculations experimental Radius and Volume of Balloo Circumference = 1000 m



 $v_{\text{balloon}} = \frac{4}{3} \pi r^3 = \frac{4}{3} (9.79 \text{ cm})^3$

Experimental mol and mass of $CO_2(g)$ O(g) O(g

P = 1.00 atm, V =

 $mass CO_2 = g CO_2$ $Molar Mass CO_2 = 44.01g/mol$

mol	1.07atm·3.93 L		
$CO_2 =$	0.0821 Latin 296.2k		
mol CO ₂ =	O.17molCO2		

mass CO ₂ =	0.1/m/CO2	44.01gCO2
		Imolico _z
$\begin{array}{c} mass \\ CO_2 = \end{array}$	7.48gCC)2