

Isotope: Form of an element based on # n°

Name _____ $^{12}_6\text{C}: 6n^{\circ}$ $^{13}_6\text{C}: 7n^{\circ}$ $^{14}_6\text{C}: 8n^{\circ}$ Period _____

College Prep Chemistry of the Earth System

Assignment 1H – Isotope Series and Atomic Mass

20 Point

Complete the following Chart based on the isotope Series No Work = No Credit

$\text{Mass } p^{+} + n^{\circ}$

Isotope Notation	Atomic #	Mass #	Protons (p^{+})	Electrons (e^{-})	Neutrons (n°)	Atomic Mass (amu)
$^{59}_{28}\text{Ni}$	28	59	28	28	$59 - 28 = 31$	59 amu
$^{40}_{18}\text{Ar}$						
$^{9}_{4}\text{Be}$	4	9	4	4	$9 - 4 = 5$	9 amu
$^{107}_{47}\text{Ag}$						
Gallium-69						
Gallium-71						
$^{35}_{17}\text{Cl}$						
$^{37}_{17}\text{Cl}$						
Potassium-39						
Potassium-40						
Potassium-41						
$^{204}_{82}\text{Pb}$						
$^{206}_{82}\text{Pb}$						
$^{207}_{82}\text{Pb}$						
$^{208}_{82}\text{Pb}$						

Atomic Mass #

Mass # in amu

$1n^{\circ} = 1\text{amu}$

$1p^{+} = 1\text{amu}$

$1e^{-} = 0\text{amu}$

Mass # $\rightarrow$
Atomic # $\rightarrow$