Name _____ Period _____

Lab 2 – Density of a Solid and a Liquid **40 Points**

Introduction

In this lab we will be solving for density, the mass (weight) of an object based on the object's volume (*size*). To solve for density we need to measure the volume and mass of the sample using basic laboratory techniques.

Density is found using the density equation:

D = m / v, where D = Density (g/mL), m = mass (g), and v = volume (mL)

Volume by Water Displacement

The volume of a solid can also be found using the displacement method. To use the displacement method the initial volume (*initial, without solid*) and final volume (*with solid*) using the equation below:

$V_{solid} = V_{after} - V_{start}$

Density Tables for Metals

Element	Aluminum	Brass	Copper	Gold	Iron	Lead	Nickle
Density	2.60g/mL	8.50g/mL	8.96g/mL	19.3g/mL	7.87g/mL	11.3g/mL	8.90g/mL

Element	Platinum	Silver	Steel	Tin	Tungsten	Vanadium	Zinc
Density	21.5g/mL	10.5g/mL	7.86g/mL	7.26g/mL	4.51g/mL	6.0g/mL	7.14g/mL

Lab Procedure

Density of a Solid

Volume by Displacement

1. Add water to the halfway mark on a graduated cylinder (*record volume*)
2. Add metal and record new volume (*record volume*)
3. Subtract new volume from original volume
4. Use volume and mass above to solve for density
5. Determine metal based on chart above/properties

Data Table with Initial Calculations (V_{metal} *and Radius*)

Metal	$V_{initial}$ (*ml*)	V_{final} (*ml*)	V_{metal} ($V_{final} - V_{initial}$)	Mass (*g*)
Dark Silver				
Light Silver				
Orange Yellow				

Calculations $D = \text{Mass (m)} / V_{metal}$

Density (*Solid Measurements*)

Silver (*Shiny*)	Silver (*Not Shiny*)	Orange/Yellow
D = _____	D = _____	D = _____
D =	D =	D =

Results

What were the metals for each metal cylinder?

Compare results from each density calculations to chart on front of lab document

Cylinder	Silver (*Shiny*)	Silver (*Not Shiny*)	Orange/Yellow
Metal			