	Name		Period
	College Prep Chemistry of th	e Earth System	
	Noteset 1C (Part 2) – In Class	*	20 Points
	Complete the following Chart to find the average atomic mass  Carbon -12 — Mass#		
	Carbon	Isotope Atomic	Frac. Abund. Ratio
	Neon has the following isotopes	Mass	Flac. Abund. Ratio
. /	<u></u>	(12   17 (100)	
9/1	Isotope Fractional Abundance		x U.999 =  1.8188  am
10 0100N 1	C-12 0.9899	12	2 2 10 0
100%		C-13 3 amu	$ x () ()  ()() _{=}  ()  3 _{(0,0)}$
100/	C-13 0.0100		0.0.00 - 0.10 MINIO
	C-14 0.0001	1/1	A A A     A A A
	Calculate the Average	C-14     Y AMV	$ \mathbf{x} (\mathbf{x} \mathbf{x})(\mathbf{x} \mathbf{x}) =  \mathbf{x} (\mathbf{x} \mathbf{x})$
	Atomic Mass of C		1 0.0001 D.0011 MMD
	D 41 C 12	D .: C 12	D. C. C. 14
	Ratio C-12	Ratio C-13	Ratio C-14 Avg Atomic C
	11 0700	12	
	11.8788 am) + C	1.15 amu  + ().1	0014  amy = 12.002  amy
			10.000 Jemy
	$\Lambda$	٨٨	
	Average Ato	mic Wass	% abund
		VIII C     1 (0.3.3.3	70 00000 .
			I some of 1:   X/
	3		Isotopes of Li
			$1 \circ 1 \cdot 6 \circ $
			3 Li 6 amu 7.59%
	Lithium	Λ ΛΙ	7 7 7 00 01
	Lau	flug. Htomic	7 Li 7 amu 92.41%
	0.19	0 00	13 L1 141110 12.11/0
	Lithium 6.94	Mass	
		$\cap$	* % of an isotope in the Iso.   total a toms in element
		UISSST TO	I w 10 01 and 12 otope IN The
		11. 1 \ 0/ \ \	
		Highest to abund.	150. I total a tom ( In element
		0	1 10 101 0 1011 5 1 0 1
			2 = 1
	17	Ralinana	35 r l 7 7 7 7 7 9
ΛΔ		DETWEEN	17 (1 1), 1) (0
Hvg. N	tomic (	25-2L	
h۸ <sup>0</sup>	-> 35 15/		3701 711 779
Masc	77.10	rammala ISA	7761 64.66/0
	. (		35C1 75.75% 37C1 24.22%