Name _____ Period _____

Freshmen Transition

Assignment 2B – Valance Electrons and Lewis Dot Structures 20 Points

Define the following terms based on ions and ion structure

Define Orbital (*s, p, d, f*)	Define Lewis Dot Structure			
What is the most common filling order for valance electrons in an atom?	How many total electrons can each of the orbitals below contain?			
	s orbital		p orbital	
	d orbital		f orbital	

Complete the chart below based on the Lewis Dot Structure for each element shown below

Lewis Dot Structure	*Lewis Dot Structure*	*Lewis Dot Structure*	*Lewis Dot Structure*	*Lewis Dot Structure*	*Lewis Dot Structure*
Na	Al	Cl	Sr	K	Se
Periodic Group #	*Periodic Group #*	*Periodic Group #*	*Periodic Group #*	*Periodic Group #*	*Periodic Group #*
Valence e⁻	*Valence e⁻*	*Valence e⁻*	*Valence e⁻*	*Valence e⁻*	*Valence e⁻*
Lewis Dot Structure	*Lewis Dot Structure*	*Lewis Dot Structure*	*Lewis Dot Structure*	*Lewis Dot Structure*	*Lewis Dot Structure*
P	Xe	Si	Ga	Ba	Br
Periodic Group #	*Periodic Group #*	*Periodic Group #*	*Periodic Group #*	*Periodic Group #*	*Periodic Group #*
Valence e⁻	*Valence e⁻*	*Valence e⁻*	*Valence e⁻*	*Valence e⁻*	*Valence e⁻*