

Name _____ Period _____

College Prep Chemistry of the Earth

Assignment 4E – Introduction to Chemical Reactions

20 Points

Define the following terms

Chemical Reaction	Yields (\rightarrow)

Reactants (<i>of reaction</i>)	Products (<i>of reaction</i>)

Name the following states of matter as part of a chemical reaction

(s)	(l)	(g)	(aq)
Types of Matter	Types of Matter	Types of Matter	Types of Matter

Complete the following chart based on energy of a reaction

Endothermic		Exothermic	
Role in Reaction	Reactant or Product	Role in Reaction	Reactant or Product

Complete the chart for the reactions below

Na + Cl ₂ \rightarrow 2NaCl + Heat		Al ₂ O ₃ • H ₂ O + Heat \rightarrow Al ₂ O ₃ + H ₂ O	
Reactants	Products	Reactants	Products
Endo or Exothermic		Endo or Exothermic	

$\text{H}_2\text{O} + \text{Heat} \rightarrow \text{H}_2 + \text{O}_2$		$\text{Ca} + \text{Zn}_2\text{S} \rightarrow \text{Zn} + \text{CaS} + \text{Heat}$	
Reactants	Products	Reactants	Products
Endo or Exothermic		Endo or Exothermic	

$\text{AgBr} + \text{Cu}(\text{NO}_3)_3 + \text{Heat} \rightarrow$ $\text{AgNO}_3 + \text{CuBr}_3$		$\text{HF} + \text{Al}(\text{OH})_3 \rightarrow \text{H}_2\text{O} + \text{AlF}_3 + \text{Heat}$	
Reactants	Products	Reactants	Products
Endo or Exothermic		Endo or Exothermic	