

Answer the following questions based on double replacement reactions

Define <i>Double Replacement Reaction</i>	Define <i>Acid Base Reaction</i>

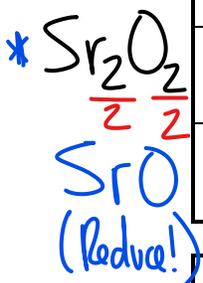
For the following reactions, write the products of each reaction following the template provided. Show work for ionic compound



Reaction	SrCl ₂ + V ₂ O ₃ → _____ + _____ V = +3	Reaction	Al ₂ O ₃ + CrF ₂ → _____ + _____ Cr = +2
AB + CD → AD + CB		AB + CD → AD + CB	
AB	SrCl ₂ CB	AB	CB
	V ₂ O ₃		

Formula	2A	A	D	6A + 3C	B	7A
Ion Charge	+2	-2	+3	-1		
Cross Method	SrO		VCl			
AD CB	SrO*		VCl ₃			

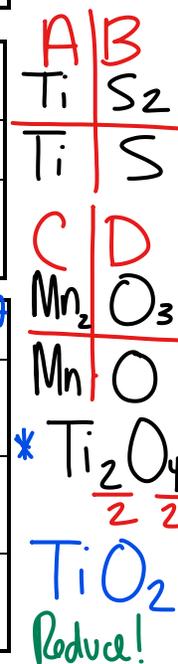
Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				



Reaction	K ₃ N + FeCl ₂ → _____ + _____ Fe = +2	Reaction	TiS ₂ + Mn ₂ O ₃ → _____ + _____ Ti = +4 Mn = +3
AB + CD → AD + CB		AB + CD → AD + CB	
AB		AB	TiS ₂ CB
			Mn ₂ O ₃

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Formula	+4	A	D	6A + 3C	B	6A
Ion Charge	+4	-2	+3	-2		
Cross Method	TiO		MnS			
AD CB	TiO ₂ *		Mn ₂ S ₃			



Reaction	$\text{AlP} + \text{CuO} \rightarrow \underline{\quad} + \underline{\quad}$ Cu = +2			Reaction	$\text{SnSe}_2 + \text{AgI} \rightarrow \underline{\quad} + \underline{\quad}$ Sn = +4 Ag = +1		
AB + CD → AD + CB				AB + CD → AD + CB			
AB		CB		AB		CB	

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Reaction	$\text{WN}_2 + \text{ZnF} \rightarrow \underline{\quad} + \underline{\quad}$ W = +6 Zn = +1			Reaction	$\text{Pb}_3\text{N}_2 + \text{FeBr}_2 \rightarrow \underline{\quad} + \underline{\quad}$ Pb = +4 Fe = +2		
AB + CD → AD + CB				AB + CD → AD + CB			
AB		CB		AB		CB	

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				