

Name \_\_\_\_\_ Period \_\_\_\_\_

College Prep Chemistry of the Earth

Assignment 4M – Writing Double Replacement Reactions w/Polyatomic Ions 20 Points

For the following reactions, write the products of each reaction following the template provided. Show work for ionic compound

Reaction	$\text{LiBr} + \text{Ca}(\text{NO}_3)_2 \rightarrow \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$	Reaction	$\text{Al}(\text{OH})_3 + \text{K}_2\text{S} \rightarrow \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$
$\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$		$\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$	
AB		CD	

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Reaction	$\text{SeCl}_2 + \text{Li}_3\text{PO}_4 \rightarrow \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$	Reaction	$\text{Mg}_3\text{N}_2 + \text{K}_2\text{Cr}_2\text{O}_7 \rightarrow \underline{\hspace{1cm}} + \underline{\hspace{1cm}}$
$\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$		$\text{AB} + \text{CD} \rightarrow \text{AD} + \text{CB}$	
AB		CD	

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Reaction	$\text{Ti}(\text{CO}_3)_2 + \text{Al}(\text{BrO}_3)_3 \rightarrow \_\_\_ + \_\_\_$ Ti = +4	Reaction	$\text{Fe}(\text{CN})_2 + \text{V}_2(\text{S}_2\text{O}_3)_3 \rightarrow \_\_\_ + \_\_\_$ Fe = +2    V = +3
AB + CD → AD + CB		AB + CD → AD + CB	
AB		CD	

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Reaction	$\text{NiC}_2\text{O}_4 + \text{Zr}_2\text{SO}_4 \rightarrow \_\_\_ + \_\_\_$ Ni = +2    Zr = +1	Reaction	$\text{Sn}(\text{NO}_3)_4 + \text{Mo}(\text{SCN})_2 \rightarrow \_\_\_ + \_\_\_$ Sn = +4    Mo = +2
AB + CD → AD + CB		AB + CD → AD + CB	
AB		CD	

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				

Formula	A	D	C	B
Ion Charge				
Cross Method				
AD CB				