

Name _____ Period _____

College Prep Chemistry of the Earth

Assignment 4S – Writing Combination Reactions with Balancing (Part 2) 20 Points

For the following reactions, write the products of each reaction following the template provided. Balance each reaction. Show work for ionic compound

Reaction	_____ Ni ³⁺ + _____ KClO ₃ → _____ + _____			
General Form	A + BC → AC + B			

Formula	A	C	Reactants		Products	
Ion Charge			Ni		Ni	
Cross Method			K		K	
AC			ClO ₃		ClO ₃	

Reaction	_____ Ba + _____ Pd(SO ₄) ₂ → _____ + _____			
General Form	A + BC → AC + B			

Formula	A	C	Reactants		Products	
Ion Charge			Ba		Ba	
Cross Method			Pd		Pd	
AC			SO ₄		SO ₄	

Reaction	_____ Cr ¹⁺ + _____ Mo ₃ (PO ₄) ₂ → _____ + _____			
General Form	A + BC → AC + B			

Formula	A	C	Reactants		Products	
Ion Charge			Cr		Cr	
Cross Method			Mo		Mo	
AC			PO ₄		PO ₄	

Reaction	$\underline{\hspace{1cm}} \text{Sc}(\text{HCO}_3)_3 + \underline{\hspace{1cm}} \text{RaF}_2 \rightarrow \underline{\hspace{1cm}} \quad + \underline{\hspace{1cm}}$ Sc = +3			
General Form	AB + CD → AD + CB			

Formula	A	D	C	B	Reactants		Products	
Ion Charge					Sc		Sc	
Cross Method					HCO ₃		HCO ₃	
AD					Ra		Ra	
CB					F		F	

Reaction	$\underline{\hspace{1cm}} \text{H}_2\text{C}_2\text{O}_4 + \underline{\hspace{1cm}} \text{Bi}(\text{OH})_4 \rightarrow \underline{\hspace{1cm}} \quad + \underline{\hspace{1cm}}$ Bi = +4			
General Form	AB + CD → AD + CB			

Formula	A	D	C	B	Reactants		Products	
Ion Charge					H		H	
Cross Method					C ₂ O ₄		C ₂ O ₄	
AD					Bi		Bi	
CB					OH		OH	

Reaction	$\underline{\hspace{1cm}} \text{Zn}_3\text{N} + \underline{\hspace{1cm}} \text{Ca}(\text{OCN})_2 \rightarrow \underline{\hspace{1cm}} \quad + \underline{\hspace{1cm}}$ Zn = +1			
General Form	AB + CD → AD + CB			

Formula	A	D	C	B	Reactants		Products	
Ion Charge					Zn		Zn	
Cross Method					N		N	
AD					Ca		Ca	
CB					OCN		OCN	