

Name _____ Period _____

College Prep Chemistry of the Earth

Assignment 5D – Molar Mass (*representative elements/compounds*)

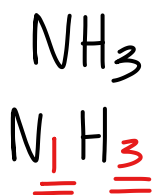
20 Points

Calculate the average molar mass for the following elements/compounds/molecules

Element	Molar Mass
Lithium (<i>Li</i>)	6.94 g/mol
Copper (<i>Cu</i>)	
Lead (<i>Pb</i>)	207.2 g/mol

Element	Molar Mass
Carbon (<i>C</i>)	
Phosphorous (<i>P</i>)	
Chlorine (<i>Cl</i>)	

For the following ionic compounds or covalent molecules, calculate the molar mass



Nitrogen Trihydride [NH_3]		
Element	#	Molar Mass
Nitrogen (N)	1	$14.01 \frac{\text{g}}{\text{mol}}$
Hydrogen (H)	3	$1.01 \frac{\text{g}}{\text{mol}}$
Nitrogen Trihydride [NH_3]		17.04 g/mol

Carbon Tetrachloride [CCl_4]		
Element	#	Molar Mass
Carbon (C)		
Chlorine (Cl)		
Carbon Tetrachloride [CCl_4]		

Sodium Fluoride [NaF]		
Element	#	Molar Mass
Sodium (Na)		
Flourine (F)		
Sodium Fluoride [NaF]		

Aluminum Oxide [Al_2O_3]		
Element	#	Molar Mass
Aluminum (Al)		
Oxygen (O)		
Aluminum Oxide [Al_2O_3]		

Lead(IV) Nitride [Pb_3N_4]		
Element	#	Molar Mass
Lead (Pb)		
Nitrogen (N)		
Lead (IV) Nitride [Pb_3N_4]		

Tungsten(VI) Sulfide [WS_3]		
Element	#	Molar Mass
Tungsten (W)		
Sulfur (S)		
Tungsten(VI) Sulfide [WS_3]		

Cobalt(II)Phosphide [Co ₃ P ₂]		
Element	#	Molar Mass
Cobalt (Co)		
Phosphorous (P)		
Cobalt(II) Phosphide [Co ₃ P ₂]		

Nitrogen Trioxide [NO ₃]		
Element	#	Molar Mass
Nitrogen (N)		
Oxygen (O)		
Nitrogen Trioxide [NO ₃]		

Aluminum Bromide [AlBr ₃]		
Element	#	Molar Mass
Aluminum (Al)		
Bromine (Br)		
Aluminum Bromide [AlBr ₃]		

Dicarbon Hexahydride [C ₂ H ₆]		
Element	#	Molar Mass
Carbon (C)		
Hydrogen (H)		
Dicarbon Hexahydride [C ₂ H ₆]		