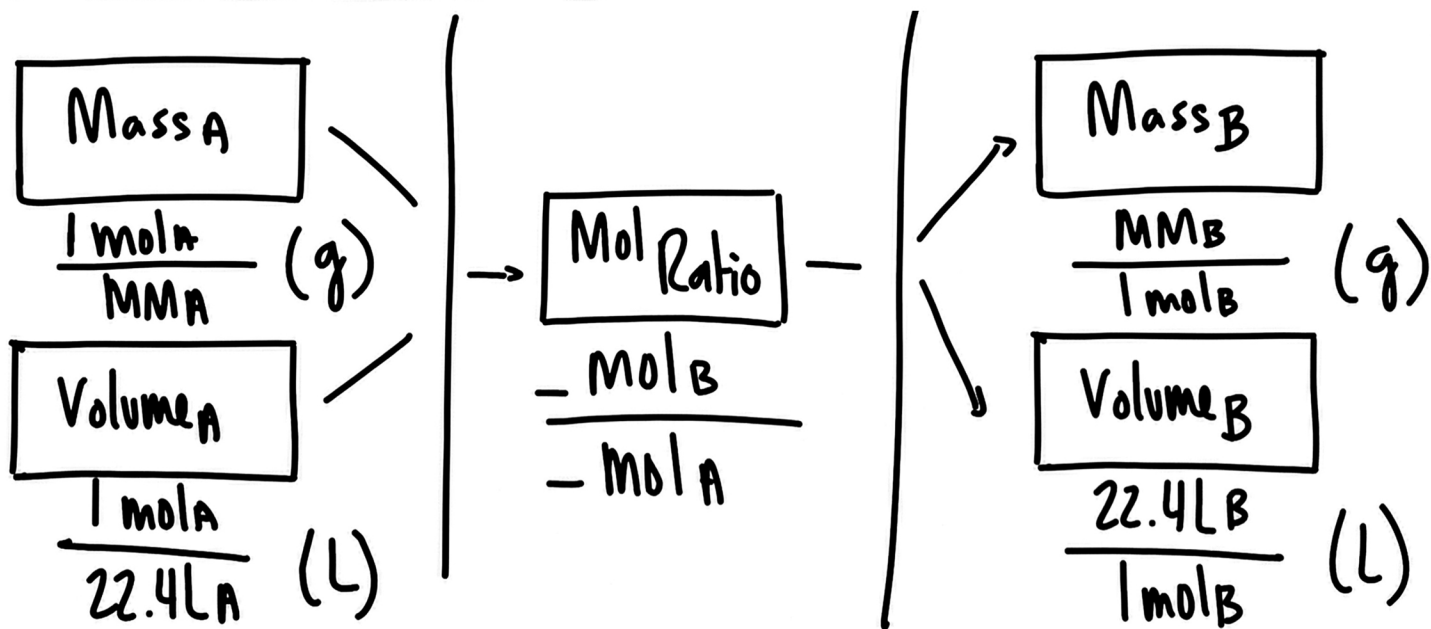


Noteset 6A (Part 5) - In Class Noteset

Three Step Conversions

Three Step Conversions



Question Forms



30.0L H_2 , convert to volume Cl_2

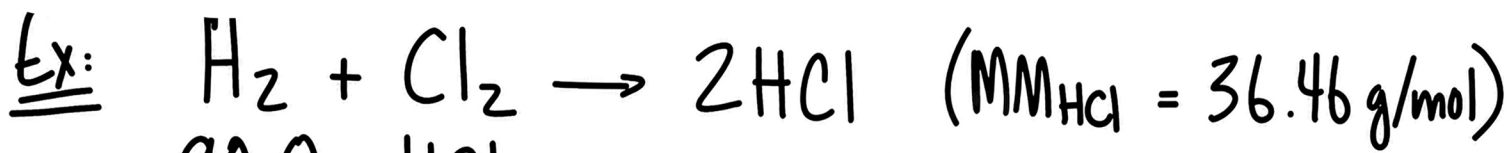
$_ \text{L } H_2 \rightarrow _ \text{mol } H_2 \quad (L_A \rightarrow L_B)$
 $_ \text{mol } H_2 \rightarrow _ \text{mol } Cl_2$
 $_ \text{mol } Cl_2 \rightarrow _ \text{volume } Cl_2$

60.0g H_2 , convert to mass HCl

$_ \text{g } H_2 \rightarrow _ \text{mol } H_2$
 $_ \text{mol } H_2 \rightarrow _ \text{mol } HCl \quad (g_A \rightarrow g_B)$
 $_ \text{mol } HCl \rightarrow _ \text{mass } HCl$

90.0g HCl , convert to volume H_2

$_ \text{g } HCl \rightarrow _ \text{mol } HCl$
 $_ \text{mol } HCl \rightarrow _ \text{mol } H_2 \quad (g_m \rightarrow L_B)$
 $_ \text{mol } H_2 \rightarrow _ \text{volume } H_2$



90.0 g HCl, convert to volume H_2

90.0g g HCl to mol HCl		— mol HCl to mol H_2		— mol H_2 to volume H_2	
90.0g HCl	1 mol HCl	2.47 mol HCl	1 mol H_2	1.24 mol H_2	22.4 L H_2
<hr/>		<hr/>		<hr/>	
	36.46g HCl		2 mol HCl		1 mol H_2
	= 2.47 mol HCl		= 1.24 mol H_2		= 34.44 L H_2